COMMON ENTRANCE TEST - 2006

AXIMUM MARKS TOTAL DURATION MAXIMUM TIME FOR ANSWERING 60 80 MINUTES 70 MINUTES MENTION YOUR CET NUMBER QUESTION BOOKLET DETAILS VERSION CODE SERIAL NUMBER MENTION YOUR CET NUMBER QUESTION BOOKLET DETAILS VERSION CODE SERIAL NUMBER Mainteend of the series of th	`CO	MMON ENTRAN	[CE TEST - 2006
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	7. Preserve the repl	ica of the OMR answer sheet fo	or a minimum period of One year.

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A -1

3

CHEMÍSTRY

1.	Which of the following is not an ore of magnesium?
	1) Carnallite2) Dolomite3) Calamine4) Sea water
2.	The atomic numbers of Ni and Cu are 28 and 29 respectively. The electron configuration
. `	$1s^2 \ 2s^2 \ 2p^6 \ 3s^2 \ 3p^6 \ 3d^{10}$ represents
	1) Cu^+ 2) Cu^{2+}
	3) Ni^{2+} 4) Ni
3.	In the following, the element with the highest ionisation energy is
	1) $[Ne]3s^2 3p^1$ 2) $[Ne]3s^2 3p^3$
	3) $[Ne]3s^2 3 p^2$ 4) $[Ne]3s^2 3 p^4$
.4.	In the conversion of Br_2 to BrO_3^- , the oxidation number of Br changes from
	1) zero to + 5 2) + 1 to + 5
•	3) $zero to - 3$ 4) $+ 2 to + 5$
5.	Among the alkali metals cesium is the most reactive because
	1) its incomplete shell is nearest to the nucleus
	2) it has a single electron in the valence shell
•	3) it is the heaviest alkali métal
· ·	4) the outermost electron is more loosely bound than the outermost electron of the

other alkali metals.

	* 4	A -1
6.	Which of the following represents the Lewis structure of N_2 molecule ?	o į
•	1) ${}^{\times}_{\times}N \equiv N^{\times}_{\times}$ 2) ${}^{\times}_{\times}N^{\times} \equiv N^{\times}_{\times}$	
	3) $\underset{x \times x}{\overset{\times}{N}} - \underset{x \times}{\overset{\times}{N}} \overset{\times}{\overset{\times}{x}}$ 4) $\underset{x \times x}{\overset{\times}{N}} \overset{\times}{\overset{\times}{x}} = \underset{x \times x}{\overset{\times}{N}} \overset{\times}{\overset{\times}{x}}$	
7.	Hydrogen bond is strongest in	
	1) $S - H O$ 2) $O - H S$ 3) $F - H F$ 4) $O - H N$	c be
8.	The decomposition of a certain mass of $CaCO_3$ gave 11.2 dm ³ of CO_2 gas at ST of KOH required to completely neutralise the gas is	P. The mass
	1) 56 g 2) 28 g 3) 42 g 4) 20 g	
9.	The density of a gas is 1.964 g dm ⁻³ at 273 k and 76 cm Hg. The gas is	् •
	1) CH_4 3) CO_2 2) C_2H_6 4) Xe	
10.	0.06 mole of KNO_3 solid is added to 100 cm ³ of water at 298 k. The enthalpy solution is 35.8 kJmol ⁻¹ . After the solute is dissolved the temperature of the solution	y of <i>KNO_{3aq}</i> ution will be
	1) 293 k 2) 298 k	
	3) 301 k 4) 304 k	

5

11. 4 moles each of SO_2 and O_2 gases are allowed to react to form SO_3 in a closed vessel. At equilibrium 25 % of O_2 is used up. The total number of moles of all the gases present at equilibrium is

- 1)
 6.5
 2)
 7.0

 3)
 8.0
 4)
 2.0
- 12. An example for autocatalysis is
 - 1) oxidation of NO to NO_2
 - 2) oxidation of SO_2 to SO_3
 - 3) decomposition of $KClO_3$ to KCl and O_2
 - 4) oxidation of oxalic acid by acidified $KMnO_4$
- 13. During the fusion of an organic compound with sodium metal, nitrogen of the compound is converted into
 - 1) NaNO₂
 2) NaNH₂
 3) NaCN
 4) NaNC
- 14. Identify the product Y in the following reaction sequence

$$CH_2 - CH_2 - COO$$

$$Ca \xrightarrow{heat} X \xrightarrow{Zn - Hg} HCl, heat$$

$$CH_2 - CH_2 - COO$$

$$(1) \quad \text{pontane}$$

$$(2)$$

- 3) cyclopentane 4) cyclopentanone
- 15. The reaction $C_2H_5ONa + C_2H_5I \rightarrow C_2H_5OC_2H_5 + NaI$ is known as
 - Kolbe's synthesis
 Williamson's synthesis
 Grignard's synthesis
 - 3) Williamson's synthesis4) Grignard's synthesis

(Space for Rough Work)

cyclobutane

 ΔG° Vs T plot in the Ellingham's diagram slopes downwards for the reaction 16. $\cdot 1) \quad Mg + \frac{1}{2}O_2 \to MgO$ 2) $2Ag + \frac{1}{2}O_2 \rightarrow Ag_2O$ 3) $C + \frac{1}{2}O_2 \rightarrow CO$ 4) $CO + \frac{1}{2}O_2 \rightarrow CO_2$ Which of the following reaction taking place in the Blast furnace is endothermic? 17. 1) $CaCO_3 \rightarrow CaO + CO_2$ 2) $2C + O_2 \rightarrow 2CO$ 3) $C + O_2 \rightarrow CO_2$ 4) $Fe_2O_3 + 3CO \rightarrow 2Fe + 3CO_2$ Liquor ammonia bottles are opened only after cooling. This is because 18. 1) it is a mild explosive 2) it is a corrosive liquid 3) it is a lachrymatory 4) it generates high vapour pressure The formation of $O_2^+[P_t F_6]^-$ is the basis for the formation of Xenon fluorides. This is because 19. 1) O_2 and Xe have comparable sizes 2) both O_2 and Xe are gases 3) O_2 and Xe have comparable ionisation energies 4) O_2 and Xe have comparable electronegativities The highest magnetic moment is shown by the transition metal ion with the configuration 20. 1) $3d^2$ 2) $3d^{5}$ 3) $3d^{7}$ 4) $3d^{9}$ (Space for Rough Work)

A -1

21.	• A transition metal ion	exists in its highest	oxidation state.	It is	expec	ted t	o behave as	
	2							

- 1) a chelating agent 2) a central metal in a coordination compound
- 3) an oxidising agent 4) a reducing agent

22. In which of the following complex ion, the central metal ion is in a state of sp^3d^2 hybridisation?

- (1) $[CoF_6]^{3-}$ (2) $[Co(NH_3)_6]^{3+}$
 - 3) $\left[Fe(CN)_{6}\right]^{3-}$ 4) $\left[Cr(NH_{3})_{6}\right]^{3+}$

23. Which of the following can participate in linkage isomerism?

1) NO_2^- 3) H_2O 2) $H_2\ddot{N}CH_2CH_2\ddot{N}H_2$ 4) : NH_3

24. Which of the following has the highest bond order ?

1) N ₂		· · · ·	2)	O_2
3) He ₂	•		4)	H_2

25. Which of the following is diamagnetic?

<u>`</u> 1)	${H_2}^+$	• • •		2)	O_2
	Li ₂	:	Ð	4)	He_2^+

	A -1
26	• The concentration of a reactant X decreases from 0.1 M to 0.025 M in 40 minutes. If the reaction follows I order kinetics, the rate of the reaction when the concentration of X is 0.01 M will be
•	1) $1.73 \times 10^{-4} M \min^{-1}$ 2) $3.47 \times 10^{-4} M \min^{-1}$
• •	3) $3.47 \times 10^{-5} M \min^{-1}$ 4) $1.73 \times 10^{-5} M \min^{-1}$
27.	Chemical reactions with very high E_a values are generally1) very fast2) very slow3) moderately fast4) spontaneous
28.	Which of the following does not conduct electricity ?
· · ·	1) fused $NaCl$ 2) solid $NaCl$
	3) brine solution 4) Copper
29.	When a quantity of electricity is passed through $CuSO_4$ solution, 0.16 g of Copper gets deposited. If the same quantity of electricity is passed through acidulated water, then the volume of H_2 liberated at STP will be [Given At.Wt. $Cu = 64$]1) 4.0 cm ³ 2) 56 cm ³ 3) 604 cm ³ 4) 8.0 cm ³
30.	Solubility product of a salt AB is 1×10^{-8} M ² in a solution in which the concentration of A ⁺ ions is 10^{-3} M. The salt will precipitate when the concentration of D ₂ :

ions is 10^{-3} M. The salt will precipitate when the concentration of B^- ions is kept

1) between 10^{-1}	⁸ M to 10 ⁻⁷	Μ	 2) between 10^{-7} M	I to 10 ⁻⁶ M
3) > 10^{-5} M		• •	4) $< 10^{-8}$ M	

4) < 10^{-8} M

9

31. Which one of the following condition will increase the voltage of the cell represented by the

equation: $Cu_{(s)} + 2Ag^{+}_{aq} \rightleftharpoons Cu^{2+}_{aq} + 2Ag_{(s)}$

- 1) increase in the dimensions of Cu electrode
- 2) increase in the dimensions of Ag electrode
- 3) increase in the concentration of Cu^{2+} ions
- 4) increase in the concentration of Ag^+ ions
- **32.** The pH of 10^{-8} M *HCl* solution is
 - 1) 8. 2) more than 8
 - 3) between 6 and 7 4) slightly more than 7
- **33.** The mass of glucose that should be dissolved in 50 g of water in order to produce the same lowering of vapour pressure as is produced by dissolving 1 g of urea in the same quantity of water is
 - 1) 1 g 3) 6 g 2) 3 g 4) 18 g
- 34. Osmotic pressure observed when benzoic acid is dissolved in benzene is less than that expected from theoretical considerations. This is because
 - 1) benzoic acid is an organic solute
 - 2) benzoic acid has higher molar mass than benzene
 - 3) benzoic acid gets associated in benzene
 - 4) benzoic acid gets dissociated in benzene
- **35.** For a reaction to be spontaneous at all temperatures
 - 1) ΔG and ΔH should be negative 2) ΔG and ΔH should be positive
 - 3) $\Delta G = \Delta S = 0$ 4) $\Delta H < \Delta G$

		, 10 A - 1
36.	. Which of the following electrolyte w	ill have maximum flocculation value for $Fe(OH)_3$ sol. ?
•	1) $NaCl$	2) $Na_{2}S$
	3) $(NH_4)_3 PO_4$	4) $K_2 SO_4$
37.	. For a reversible reaction : $X_{(g)} + 3Y_{(g)}$	$_{(g)} \rightleftharpoons 2Z_{(g)}$
•	$\Delta H = -40 \text{kJ}$ the standard entropies The temperature at which the above	s of X, Y and Z are 60, 40 and 50 JK ⁻¹ mol ⁻¹ respectively. The reaction attains equilibrium is about
	1) 400 K	2) 500 K
	3) 273 K	4) 373 K
38.	The radii of Na^+ and Cl^- ions are 95 p unit cell is	pm and 181 pm respectively. The edgé length of $NaCl$
• .	1) 276 pm	2) 138 pm
• •	3) 552 pm	4) 415 pm
39.	Inductive effect involves	
	1) displacement of σ electrons	s 2) delocalisation of π electrons
- 1 ^{- 1} -	3) delocalisation of σ electrons	
40.	The basicity of aniline is less than tha	
•.	1) D 00	
	3) $-R$ effect of $-NH_2$ group	 2) - I effect of - NH₂ group 4) hyperconjugation effect

•••		A-1
41	Methyl bromide is converted into ethan	ne by heating it in ether medium with
	1) <i>Al</i>	$2) \Delta n$
	3) <i>Na</i>	4) <i>Cu</i>
42.	Which of the following compound is exp	pected to be optically active ?
• • •	1) $(CH_3)_2 CH CHO$	2) $CH_3CH_2CH_2CHO$
	3) CH ₃ CH ₂ CHBr CHO	4) $CH_3CH_2CBr_2CHO$
43.	Which cycloalkane has the lowest heat	t of combustion per $C\!H_2$ group ?
40.	1) cyclopropane	2) cyclobutane
	3) cyclopentane	4) cyclohexane
44.	The catalyst used in the preparation alcohol is	of an alkyl chloride by the action of dry <i>HCl</i> on an
	1) anhydrous $AlCl_3$	2) <i>FeCl</i> ₃
	3) anhydrous $ZnCl_2$	4) <i>Cu</i>
45.		
. •	$R - X \xrightarrow{alcoholic} KCN \to A \xrightarrow{dilute} B,$	
	the product B is	0) ll-bla
	1) alkyl chloride	2) aldehyde4) ketone
•	3) carboxylic acid	4) ketone
<u> </u>	(Space	e for Rough Work)

Turn Over

		.12	A -1
46.	Which of the following compound way		2
	Which of the following compound wou 1) salicylic acid	Id not evolve CO_2 when treated with	$NaHCO_3$ solution?
· - ·	3) benzoic acid	 2) phenol 4) 4-nitro benzoic acid 	
47.	By heating phenol with chloroform in		
	1) salicylic acid	2) salicylaldehyde	٥
•	3) anisole	4) phenyl benzoate	<i>1</i> ,
48.	When a mixture of calcium benzoat compound is		illed, the resulting
	1) acetophenone	2) benzaldehyde	η.
'. ·	3) benzophenone	4) acetaldehyde	
49.	Which of the following does not give b	benzoic acid on hydrolysis ?	
	1) phenyl cyanide	2) benzoyl chloride	· · · · · · · · · · · · · · · · · · ·
	3) benzyl chloride	4) methyl benzoate	
50.	Which of the following would undergo		
	0	, Termania reaction to Bive a prime	ary amine :
• • •			
· · · ·	1) $R-C-Cl$	2) $RCONHCH_3$	

A -1

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51. Glucose contains in addition to aldehyde group

- 1) one secondary OH and four primary OH groups
- 2) one primary OH and four secondary OH groups
- 3) two primary OH and three secondary OH groups
- 4) three primary OH and two secondary OH groups
- 52. A distinctive and characteristic functional group of fats is
 - 1) a peptide group 2) an ester group
 - 3) an alcoholic group 4) a ketonic group
- **53.** At pH = 4 glycine exists as
 - 1) $H_3 \overset{+}{N} CH_2 COO^-$ 2) $H_3 \overset{+}{N} CH_2 COOH$ 3) $H_2 N CH_2 COOH$ 4) $H_2 N CH_2 COO^-$
- 54. Insulin regulates the metabolism of

1) minerals	· · . ·	· •		.2)	amino acids
I) immerans			·		

3) glucose 4) vitamins

55. The formula mass of Mohr's salt is 392. The iron present in it is oxidised by $KMnO_4$ in acid medium. The equivalent mass of Mohr's salt is

1)	392	•	·		•••	-	. ·	2)	31.6
	278	•		•			• • • •	4)	156

A -1

1) the reduction of nitrate to nitric oxide oxidation of nitric oxide to nitrogen dioxide 2) reduction of ferrous sulphate to iron 3) oxidising action of sulphuric acid 4) Acrolein test is positive for 57. 1) polysaccharides 2) proteins 3) oils and fats 4) reducing sugars An organic compound which produces a bluish green coloured flame on heating in presence **58.** of copper is 1) chlorobenzene 2) benzaldehyde 3) aniline 4) benzoic acid For a reaction $A + B \rightarrow C + D$ if the concentration of A is doubled without altering the **59.**concentration of B, the rate gets doubled. If the concentration of B is increased by nine times without altering the concentration of A, the rate gets tripled. The order of the reaction is 1) 2 2) (4) $\frac{4}{3}$ 3) Which of the following solutions will exhibit highest boiling point? 60. $0.01 \text{ M } Na_2 SO_4_{(aq)}$ 0.01 M KNO_{3 (aq)} 1) 2) $0.015 \text{ M urea}_{(aq)}$ 4) 0.015 M glucose_(aq) 3) (Space for Rough Work)

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The brown ring test for nitrates depends on

56.



A - 1

સરપ્રદેશ કર્મા કરી હત્ત કરી

SR - 17

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