



Rao IIT Academy

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JEE | MEDICAL-UG | BOARDS | KVPY | NTSE | OLYMPIADS

XII - ISC Board

Date: 05.03.2018

CHEMISTRY - QP + SOLUTIONS

Question 1

(a) Fill in the blanks by choosing the appropriate word/words from those given in the brackets:

[4×1]

(square pyramidal, electrical, 74; 26, sp^3d^2 , sp^3d , chemical, 68, 32, tetrahedral, yellow, white, iodoform, Lucas)

(i) A Galvanic cell converts _____ energy into _____ energy.

Ans. A Galvanic cell converts **chemical** energy into **electrical** energy.

Topic: Electrochemistry_Subtopic: Galvanic cell_Level: Easy_ISC Board / Chemistry

(ii) The percentage of unoccupied spaces in bcc and fcc arrangements are _____ and _____ respectively.

Ans. The percentage of unoccupied spaces in bcc and fcc arrangements are **32** and **26** respectively.

Topic: Solid state_Subtopic: Packing_Level: Easy_ISC Board / Chemistry

(iii) Propan-2-ol on reaction with iodine and sodium hydroxide gives _____ precipitate and the reaction is called _____ test.

Ans. Propan-2-ol on reaction with iodine and sodium hydroxide gives **yellow** precipitate and the reaction is called **iodoform** test.

Topic: Alcohol_Subtopic: Iodoform test_Level: Easy_ISC Board / Chemistry

(iv) The geometry of $XeOF_4$ molecule is _____ and the hybridisation of Xenon atom in the molecule is _____.

Ans. The geometry of $XeOF_4$ molecule is **pyramidal** and the hybridisation of Xenon atom in the molecule is **sp^3d^2** .

Topic: p-block_Subtopic: Group-18_Level: Medium_ISC Board / Chemistry

(b) Complete the following statements by selecting the correct alternative from the choices given: [4×1]

(i) During the course of an S_N1 reaction, the intermediate species formed is:

(1) a carbocation (2) a free radical (3) a carbanion (4) an intermediate complex

Ans. (1) a carbocation

Topic: HDA_Subtopic: S_N1 _Level: Easy_ISC Board / Chemistry

(ii) Purification of aluminium by electrolytic refining is called:

(1) Serpeck's process (2) Hoope's process

(3) Hall's process (4) Baeyer's process

Ans. (2) Hoope's process

Topic: General principles_Subtopic: Extraction of Al_Level: Easy_ISC Board / Chemistry

- (iii) An aqueous solution of urea freezes at -0.186°C , K_f for water = $1.86 \text{ K kg mol}^{-1}$, K_b for water = $0.512 \text{ K kg mol}^{-1}$. The boiling point of urea solution will be:
 (1) 373.065 K (2) 373.186 K (3) 373.512 K (4) 373.0512 K

Ans. (4) 373.0512 K

$$[0 - (-0.186)] = 1.86 \times m$$

$$\therefore m = \frac{0.186}{1.86} = 0.1$$

$$\Delta T_b = K_b \cdot m$$

$$\therefore \Delta T_b = 0.512 \times 0.1$$

$$= 0.0512$$

$$T - T^{\circ} = 0.0512$$

$$T = 0.0512 + T^{\circ}$$

$$= 0.0512 + 373$$

$$= 373.0512 \text{ K}$$

Topic: Solutions_Subtopic: Depression of freezing point_Level: Medium_ISC Board / Chemistry

- (iv) In the dehydration of alcohols to alkenes by heating with concentrated sulphuric acid, the initiation step is:

- (1) formation of carbocation (2) formation of an ester
 (3) protonation of alcohol molecule (4) elimination of water

Ans. (3) protonation of alcohol molecule

Topic: Alcohols_Subtopic: Dehydration_Level: Easy_ISC Board / Chemistry

- (c) Match the following: [4×1]

- | | |
|-------------------------------|------------------------|
| (i) Rate constant | (a) Dialysis |
| (ii) Biodegradable polymer | (b) Glycine |
| (iii) Zwitter ion | (c) Arrhenius equation |
| (iv) Purification of colloids | (d) PHBV |

Ans. (i) Rate constant Arrhenius equation
 (ii) Biodegradable polymer PHBV
 (iii) Zwitter ion Glycine
 (iv) Purification of colloids Dialysis

Topic: Mixed_Subtopic: Mixed_Level: Easy_ISC Board / Chemistry

- (d) Answer the following questions: [4×2]

- (i) (1) Why does the density of transition elements increase from Titanium to Copper? (at. no. Ti = 22, Cu = 29)

Ans. Across the period from Ti to Cu as atomic size decreases, the density of element increases.

Topic: d-block_Subtopic: Density_Level: Easy_ISC Board / Chemistry

- (2) Why is zinc not regarded as a transition element?
 (at. no. Zn = 30)

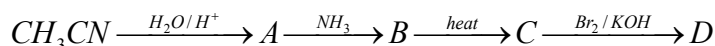
Ans. $\text{Zn}(30) - [\text{Ar}]_{18} 4s^2 3d^{10}$

Transition metals show vacant d-orbitals but Zn metal has completely filled d-orbitals.

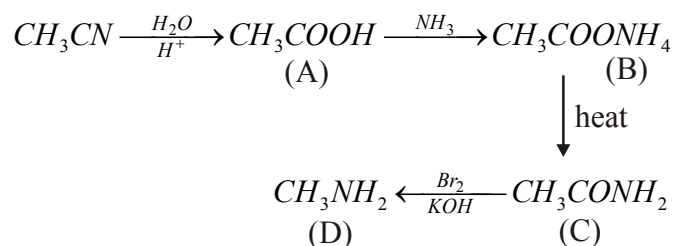
\therefore Zn is not regarded as transition element.

Topic: d-block_Subtopic: Properties_Level: Easy_ISC Board / Chemistry

(ii) Identify the compounds A, B, C and D.



Ans.



Topic: Carboxylic acids **Subtopic:** Mixed **Level:** Difficult **ISC Board / Chemistry**

(iii) Calculate the osmotic pressure of a solution prepared by dissolving 0.025 g of K_2SO_4 in 2.0 litres of water at 25°C assuming that K_2SO_4 is completely dissociated. (mol. wt. of $\text{K}_2\text{SO}_4 = 174 \text{ g mol}^{-1}$)

Ans. Mass of $\text{K}_2\text{SO}_4 = 0.025 \text{ gm}$
 Molar mass of $\text{K}_2\text{SO}_4 = 174 \text{ g/mol}$
 Volume = 2 L
 Temperature = 25°C (298 K)

$$\pi V = nRT$$

$$\pi = \frac{n}{V} RT$$

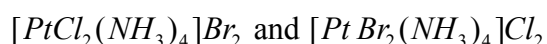
$$= \frac{\text{mass}}{\text{molar mass}} \times \frac{1}{V} RT$$

$$= \frac{0.025}{174} \times \frac{1}{2} \times 8.314 \times 298$$

$$= 0.173 \text{ atm}$$

Topic: Solutions **Subtopic:** Osmotic pressure **Level:** Medium **ISC Board / Chemistry**

(iv) What type of isomerism is shown by the following coordination compounds:



Write their IUPAC names.

Ans. These coordination compounds show ionization isomerism.

$[\text{PtCl}_2(\text{NH}_3)_4]\text{Br}_2$: Tetraammine dichloro platinum (IV) bromide

$[\text{PtBr}_2(\text{NH}_3)_4]\text{Cl}_2$: Tetraammine dibromo platinum (IV) chloride

Topic: Co-ordination compounds **Subtopic:** IUPAC **Level:** Easy **ISC Board / Chemistry**

Question 2

(a)

(i) Write the rate law expression for the reaction $A + B + C \rightarrow D + E$, if the order of reaction is first, second and zero with respect to A, B and C, respectively.

Ans. Rate law expression

$$\text{Rate} = K[A]^1.[B]^2[C]^0$$

Topic: Chemical kinetics **Subtopic:** Order **Level:** Easy **ISC Board / Chemistry**

- (ii) How many times the rate of reaction will increase if the concentration of A , B and C are doubled in the equation given in (i) above?

Ans. If concentration of A is doubled then rate also increases by double. If conc. of B is doubled then rate increases by four times and conc. of C is doubled then no change in rate of reaction.

Topic: Chemical kinetics_Subtopic: Order_Level: Easy_ISC Board / Chemistry

OR

- (b) The rate of reaction becomes four times when the temperature changes from 293 K to 313 K. Calculate the energy of activation (E_a) of the reaction assuming that it does not change with temperature. ($R = 8.314 \text{ J K}^{-1} \text{ mol}^{-1}$)

Ans. $T_1 = 293 \text{ K}$ $T_2 = 313 \text{ K}$ $R = 8.314$
 $K_1 = K$ $K_2 = 4K$

$$\log \frac{K_2}{K_1} = \frac{E_a}{2.303 R} \left(\frac{T_2 - T_1}{T_1 T_2} \right)$$

$$\log \frac{4K}{K} = \frac{E_a}{2.303 \times 8.314} \left(\frac{313 - 293}{293 \times 313} \right)$$

$$0.6020 = \frac{E_a}{2.303 \times 8.314} \left(\frac{20}{293 \times 313} \right)$$

$$E_a = \frac{0.6020 \times 2.303 \times 8.14 \times 293 \times 313}{20}$$

$$= 52854 \text{ J}$$

$$= 52.854 \text{ KJ}$$

Topic: Chemical kinetics_Subtopic: Energy of activation_Level: Medium_ISC Board / Chemistry

Question 3

- (a) How do antiseptics differ from disinfectants?

Ans. Antiseptic is used to clean wounds and kill the microorganisms of living tissues.

Disinfections are not safe to be applied on living tissues. They are used to clean drains, toilets and public sanitation.

Topic: Chemistry in everyday life_Subtopic: Antimicrobials_Level: Easy_ISC Board / Chemistry

- (b) State the role of the following chemicals in the food industry:

(i) Sodium benzoate

(ii) Aspartame

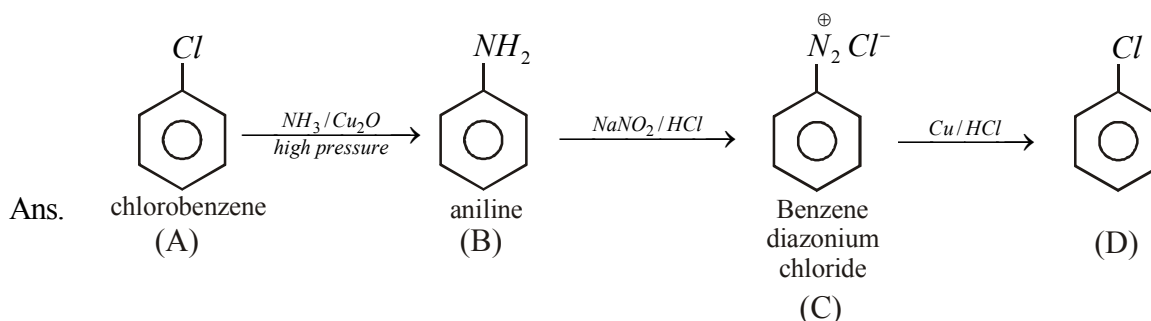
Ans. (i) Sodium benzoate : Chemical preservative

(ii) Aspartame : Artificial sweetener

Topic: Chemistry in everyday life_Subtopic: Food_Level: Easy_ISC Board / Chemistry

Question 4

An aromatic organic compound [A] on heating with NH_3 and Cu_2O at high pressure gives [B]. The compound [B] on treatment with ice cold solution of NaNO_2 and HCl gives [C], which on heating with Cu/HCl gives compound [A] again. Identify the compounds [A], [B] and [C]. Write the name of the reaction for the conversion of [B] to [C].



Conversion of B to C is called Sand-Meyer reaction.

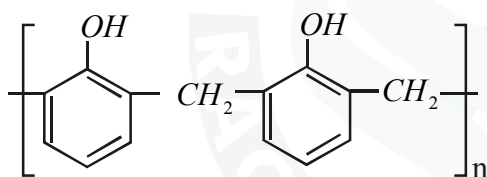
Topic: Mixed **Subtopic:** Mixed **Level:** Medium **ISC Board / Chemistry**

Question 5

Write the names of the monomers for each of the following polymers:

- (a) Bakelite
(b) Nylon-2-nylon-6

Ans. (a) Monomer of Bakelite



Novolac

Its, a phenol formaldehyde resin

- (b) Monomer of Nylon-2-nylon-6 is glycine and amino caproic acid.

Topic: Polymers **Subtopic:** Polymers **Level:** Medium **ISC Board / Chemistry**

Question 6

Name the purine bases and pyrimidine bases present in RNA and DNA.

Ans. In DNA: Adenine, Guanine, Cytosine, Thymine

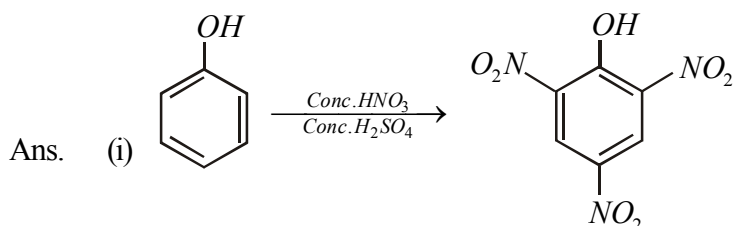
RNA: Adenine, Guanine, Cytosine, Uracil.

Topic: Biomolecules **Subtopic:** Nucleic acid **Level:** Easy **ISC Board / Chemistry**

Question 7

(a) How will you obtain the following? (Give balanced equation.)

- (i) Picric acid from phenol
(ii) Ethyl chloride from diethyl ether.



- (ii) $\text{CH}_3\text{CH}_2 - \text{O} - \text{CH}_2 - \text{CH}_3 + \text{PCl}_5 \xrightarrow{\Delta} 2\text{CH}_3\text{CH}_2 - \text{Cl} + \text{POCl}_3$

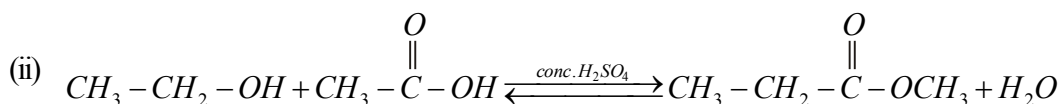
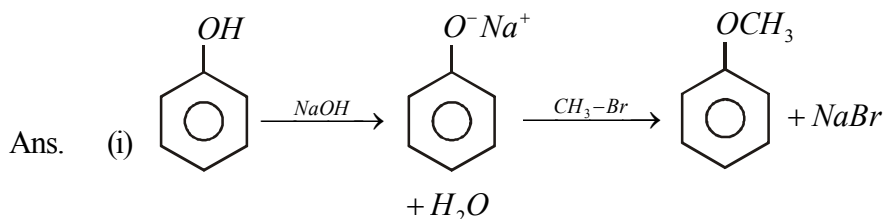
Topic: Alcohols, Phenols, Ethers **Subtopic:** Phenols and ethers **Level:** Easy **ISC Board / Chemistry**

OR

(b) How will you obtain the following? (Give balanced equation.)

(i) Anisole from phenol

(ii) Ethyl acetate from ethanol.

**Topic: Mixed_Subtopic: Mixed_Level: Medium_ISC Board / Chemistry****Question 8**

40% of a first order reaction is completed in 50 minutes. How much time will it take for the completion of 80% of this reaction?

Ans. The integrated rate Equation for 1st order reaction, is

$$K = \frac{2.303}{t} \log \frac{[A]_0}{[A]_t}$$

According to question.

$$K = \frac{2.303}{50 \text{ min}} \log \frac{[100]}{[60]}$$

$$\Rightarrow K = 0.0102 \text{ min}^{-1}$$

$$K = \frac{2.303}{t} \log \frac{100}{20}$$

$$\Rightarrow t = \frac{2.303}{0.0102} \log 5 \text{ min}$$

$$\Rightarrow t = 157.81 \text{ min}$$

Topic: Chemical kinetics_Subtopic: Integrated rate law_Level: Medium_ISC Board / Chemistry**Question 9**(a) The freezing point of a solution containing 5.85 g of NaCl in 100 g of water is -3.348°C . Calculate van't Hoff factor for this solution. What will be the experimental molecular weight of NaCl? (K_f for water = $1.86 \text{ K kg mol}^{-1}$, at wt, Na = 23, Cl = 35.5)

Ans.

$$\text{Molality } (m) = \frac{\text{no. of mole of solute}}{\text{mass of solvent in kg}}$$

$$\text{no. of mole of NaCl} = \frac{5.85}{58.5} = 0.1$$

$$m = \frac{0.1 \text{ mole}}{0.1 \text{ kg}}$$

$$= 1 \text{ mole/kg}$$

$$\Delta T_f = i \times k_f \times m$$

$$\Rightarrow 3.348 = i \times 1.86 \times 1$$

$$\Rightarrow i = 1.8$$

$$i = \frac{\text{Theoretical molar mass}}{\text{Experimental molar mass}}$$

$$\Rightarrow 1.8 = \frac{58.5}{\text{Experimental molar mass}}$$

$$\Rightarrow \text{Experimental molar mass} = \frac{58.5}{1.8} = 32.5 \text{ gm/mol}$$

Topic: Solutions_Subtopic: Van't Hoff factor_Level: Medium_ISC Board / Chemistry

OR

- (b) An aqueous solution containing 12.48 g of barium chloride (BaCl_2) in 1000g of water, boils at 100.0832°C . Calculate the degree of dissociation of barium chloride. (K_b for water = $0.52 \text{ K kg mol}^{-1}$, at. wt. Ba = 137, Cl = 35.5)

Ans. $\Delta T_b = i \times k_b \times m$

$$\begin{aligned} \text{Molar Mass of BaCl}_2 &= 137 + 35.5 \times 2 \\ &= 137 + 70 \\ &= 207 \end{aligned}$$

$$\begin{aligned} m_2 &= \frac{12.48}{207} \\ &= 0.06 \end{aligned}$$

$$m = \frac{0.06 \text{ mole}}{1 \text{ kg}}$$

$$= 0.06 \text{ moles/kg}$$

$$\Delta T_b = i \times k_b \times m$$

$$\Rightarrow 0.0832 = i \times 0.52 \times 0.06$$

$$\Rightarrow i = 2.66$$

$$\alpha = \frac{i-1}{n-1} \quad \text{for BaCl}_2 \quad m = 1 + 2 = 3$$

$$\Rightarrow \alpha = \frac{2.66-1}{3-1}$$

$$\Rightarrow \alpha = \frac{1.66}{2} = 0.83$$

Topic: Solutions_Subtopic: Van't Hoff factor_Level: Medium_ISC Board / Chemistry

Question 10

Examine the defective crystal given below and answer the question that follows :

| | | | | |
|-------|-------|-------|-------|-------|
| A^+ | B^- | A^+ | B^- | A^+ |
| B^- | | B^- | A^+ | B^- |
| A^+ | B^- | A^+ | | A^+ |
| B^- | A^+ | B^- | A^+ | B^- |

State if the above defect is stoichiometric or non-stoichiometric. How does this defect affect the density of the crystal? Also, write the term used for this type of defect.

Ans. The above defect is stoichiometric.

As the number of ions are missing from the crystal, density decreases.

This defect is known as Schottky defect.

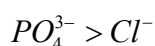
Topic: Solid state_Subtopic: Defects in crystal_Level: Medium_ISC Board / Chemistry

Question 11

Give reason for each of the following :

- For ferric hydroxide sol the coagulating power of phosphate ion is more than chloride ion.
- Medicines are more effective in their colloidal form.
- Gelatin is added to ice creams.

Ans. (a) According to Schulze-Hardy rule, to coagulate a positively charged colloid system, higher the negative charge higher will be its effectiveness.



Topic: Surface chemistry_Subtopic: Colloid_Level: Easy_ISC Board / Chemistry

(b) Medicines are more effective in their colloidal form as absorption will be better.

Topic: Surface chemistry_Subtopic: Colloid_Level: Easy_ISC Board / Chemistry

- (i) It acts as a stabilizer in ice-cream to produce smooth and creamy texture.
- (ii) It also helps to reduce ice and lactose crystal growth during storage.

Topic: Surface chemistry_Subtopic: Colloid_Level: Easy_ISC Board / Chemistry

Question 12

(a) For the complex ion $[Fe(CN)_6]^{3-}$, state :

- the type of hybridisation
- the magnetic behaviour
- the oxidation number of the central metal atom

Ans. (i) The type of hybridisation is d^2sp^3 .

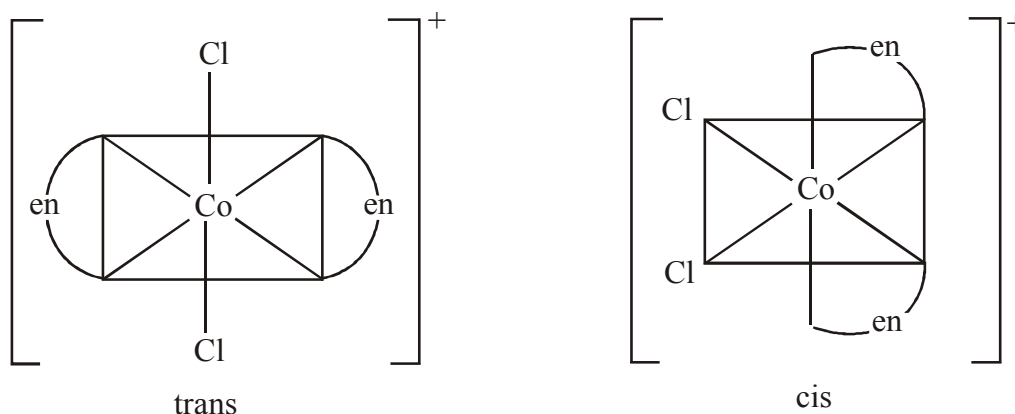
(ii) As CN^\ominus is a strong field ligand there will be only 1 unpaired electron in t_{2g} orbital so magnetic behaviour is paramagnetic.

(iii) The oxidation number of Fe is (+3) in this complex.

Topic: Coordination compounds_Subtopic: Bonding_Level: Medium_ISC Board / Chemistry

(b) Write the IUPAC name of $[Co(en)_2Cl_2]^+$ ion and draw the structures of its geometrical isomers.

Ans. IUPAC name - Dichlorobis(ethyldiammine)cobalt(III) ion.



Topic: Coordination compounds_Subtopic: Isomerism_Level: Medium_ISC Board / Chemistry

Question 13

(a) Explain why:

(i) Mn^{2+} is more stable than Fe^{2+} towards oxidation to +3 state.

(At. no. of $Mn = 25$, $Fe = 26$)

(ii) Transition elements usually form coloured ions

(iii) Zr and Hf exhibit similar properties

(At. no. of $Zr = 40$, $Hf = 72$)

Ans. (i) $Mn : [Ar] 4s^2 3d^5$ $Fe : [Ar] 4s^2 3d^6$

$Mn^{2+} : [Ar] 4s^0 3d^5$ $Fe^{2+} : [Ar] 4s^0 3d^6$

Mn^{2+} has a stable half filled d-subshell. Whereas Fe^{2+} has a partially filled d-subshell hence as Fe^{2+} loses an electron and gets converted to Fe^{3+} it achieves stable half filled d-subshell. Therefore Mn^{2+} is more stable than Fe^{2+} towards oxidation to +3 state.

(ii) In case of the transition metal ions the electrons can be easily promoted from one energy level to another in the same d-subshell. These are called d-d- transitions. This transition is responsible for corresponding colours.

(iii) (1) Zr and Hf have same number of electrons in the outermost shell.

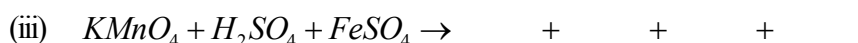
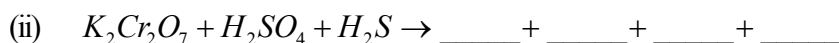
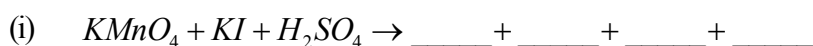
(2) They have the comparable atomic size due to lanthanoid contraction. they are called chemical twins.

Hence they exhibit same chemical properties.

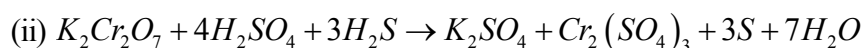
Topic: d & f block_Subtopic : Properties of transition metal_Level: Medium_ISC Board / Chemistry

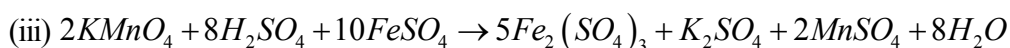
OR

Complete and balance the following chemical equations :



Ans. (i) $2KMnO_4 + 10KI + 8H_2SO_4 \rightarrow 6K_2SO_4 + 2MnSO_4 + 5I_2 + 8H_2O$ **3M**





Topic: d & f block **Subtopic :** Potassium permanganate **Level:** Medium **ISC Board / Chemistry**

Question 14

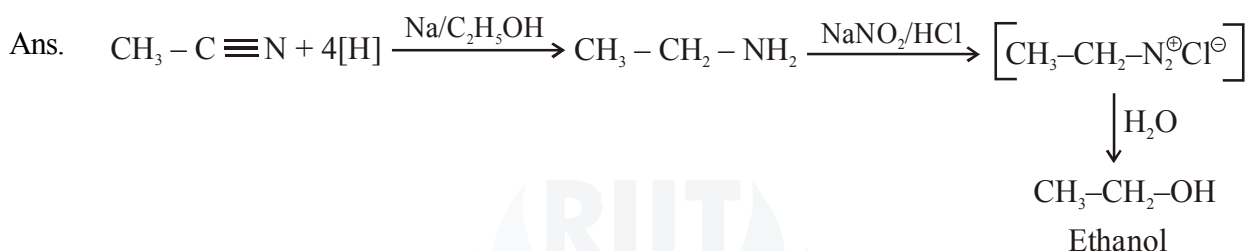
(a) Arrange the following in the increasing order of their basic strength :



Ans. $(C_2H_5)_2NH > C_2H_5NH_2 > C_6H_5NH_2$

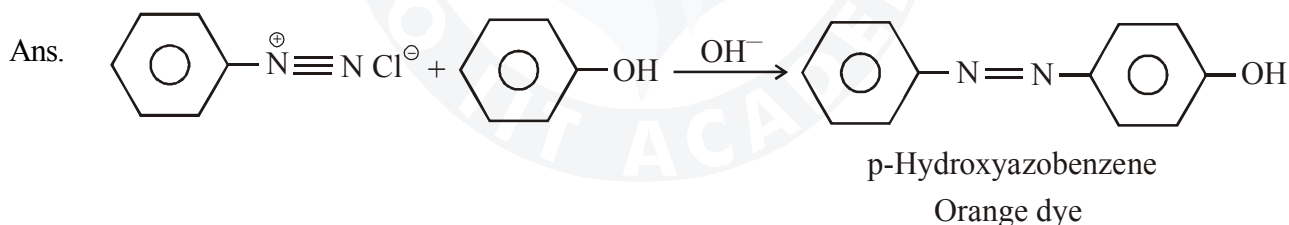
Topic: Compound containing nitrogen **Subtopic:** Basicity of Amines **Level:** Easy **ISC Board / Chemistry**

(b) Give a balanced chemical equation of convert methyl cyanide to ethyl alcohol.



Topic: Compound containing nitrogen **Subtopic:** Reactions **Level:** Easy **ISC Board / Chemistry**

(c) What happens when benzene diazonium chloride reacts with phenol in weak alkaline medium ?
(Give balanced equation).



It gives a orange coloured dye

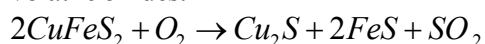
Topic: Compound containing nitrogen **Subtopic:** Diazonium compounds **Level:** Easy **ISC Board / Chemistry**

Question 15

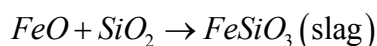
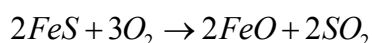
Name the sulphide ore of Copper. Describe how pure copper is extracted from this ore.

Ans. * Concentration : Copper is extracted mainly from copper pyrites. The powdered ore is concentrated by froth floatation process.

* Roasting : The concentrated ore is roasted in a limited supply of air in reverberatory furnace when a mixture of FeS and Cu_2S is formed while As and Sb present in impurities are removed as volatile oxides.



* Smelting : The roasted ore is mixed with sand and coke and heated strongly in a blast furnace when most of the iron is converted into iron oxide. It then combines with silica and is removed as slag.



The molten material thus obtained from the blast furnace consists of sulphides of Cu^+ , Fe^{2+} coke and sand. It is called matter.

- * Bessemerization : The molten mate is then transferred to a Bassemer converter and blast of air is blown. During this process, some Cu_2S is oxidized to Cu_2O which then reacts with more Cu_2S to give copper metal (auto reduction).



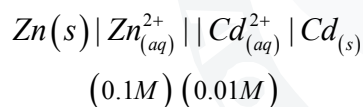
After the reductoin is complete, the molten copper is poured into sand moulds. As the metal solidifies, SO_2 escapes and leaves blisters on the surface. The solid metal thus obtained is called the blister copper and is about 99% pure.

- * Refining : Blister copper impurities of Ag and Au mainly. Impure copper is further refined electrolytically to obtain 99.95 to 99.99% pure copper by using impure copper as a anode, a rod of pure copper as cathode and acidified solutoin acid present in the electrolyte and thus settle down as anode mud while impurities of Fe, Ni, Zn ets. pass into electrolytic solution. Copper thus obtained is 99.99% pure and gets deposited at the cathode.

Topic: General principles and techniques_Subtopic:Extraction of copper_Level: Medium_ISC Board / Chemistry

Question 16

- (a) (i) Calculate the emf and ΔG° for the cell reaction at $25^\circ C$:

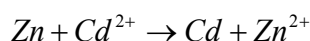


Given $E^\circ Zn^{2+} / Zn = -0.763$ and $E^\circ Cd^{2+} / Cd = -0.403V$

- (ii) Define the following terms :

- (1) Equivalent conductivity
- (2) Corrosion of metals

Ans. (i) $E_{cell}^0 = E_{Cd^{2+}/Cd}^0 - E_{Zn^{2+}/Zn}^0$
 $= -0.403 - (-0.763)$
 $= 0.360$



$$E_{cell}^0 = 0.36V$$

$$E_{cell} = E_{cell}^0 - \frac{0.0591}{2} \log \frac{[Zn^{2+}]}{[Cd^{2+}]}$$

$$= 0.36 - \frac{0.0591}{2} \log \frac{0.1}{0.01}$$

$$= 0.36 - \frac{0.591}{2} \log 10$$

$$\begin{aligned}
 &= 0.36 - \frac{0.0591}{2} \\
 &= 0.36 - 0.0295 \\
 E_{cell} &= 0.331V \\
 \Delta G^0 &= -nFE^0 \\
 &= -2 \times 96500 \times 0.36 \\
 &= -69,480J \\
 &= -69.48kJ
 \end{aligned}$$

Topic:Electrochemistry Subtopic : Cell potential Level: Easy ISC Board / Chemistry

- (ii) (a) It is defined as conductivity power of all ions produced by dissolving one gram equivalent of an electrolyte in solution. $\wedge_e = \frac{k \times 1000}{C}$

Topic:Electrochemistry Subtopic : Conductivity Level: Easy ISC Board / Chemistry

- (b) When metals are exposed to atmospheric conditions, they react with air or water in the environment to form undesirable oxides. This process is called corrosion.

Topic:Electrochemistry Subtopic : Corrosion Level: Easy ISC Board / Chemistry

- (b) (i) The specific conductivity of a solution containing 5 g of anhydrous $BaCl_2$ (mol. wt = 208) in 1000 cm^3 of a solution is found to be $0.0058 \text{ ohm}^{-1} \text{ cm}^{-1}$. Calculate the molar and equivalent conductivity of the solution.
- (ii) What is an electrochemical series? How is it useful in predicting whether a metal can liberate hydrogen from acid or not?

Ans. (i)

$$M = \frac{5}{208} \times \frac{1}{1000} \times 1000 = 0.024M$$

$$K = 0.0058$$

$$\wedge_M = \frac{1000K}{C} = \frac{1000 \times 0.0058}{0.024} = 241.66 \text{ Scm}^2 \text{ mol}^{-1}$$

$$N = \frac{5}{104} \times \frac{1}{1000} \times 1000 = 0.048N$$

$$\wedge_{eq} = \frac{1000K}{C} = \frac{1000 \times 0.0058}{0.048} = \frac{5.8}{0.048} = 120.83 \text{ Scm}^2 \text{ g.eq}^{-1}$$

Topic:Electrochemistry Subtopic : Conductivity Level: Medium ISC Board / Chemistry

- (ii) 2 M

The arrangement of elements in order of increasing reduction potential values is called electrochemical series.

Only the metals can liberate hydrogen from the acid which have negative values of reduction potential i.e. $-E^0$ values.

Topic:Electrochemistry Subtopic : Electro chemical series Level: Easy ISC Board / Chemistry

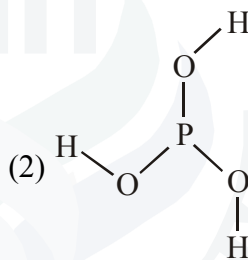
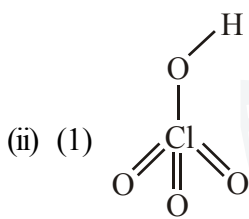
Question 17

- (a) (i) Explain why:
- (1) Nitrogen does not form pentahalides
 - (2) Helium is used for filling weather balloons
 - (3) ICl is more reactive than I_2

(ii) Draw the structures of the following :



- Ans. (i) (1) (a) Nitrogen has maximum covalency of four.
 (b) Nitrogen does not have d-orbitals.
 (c) Hence it cannot expand its octet and form hypervalent compounds.
 (2) (a) Helium is non-combustible gas.
 (b) It is lighter than air.
 (c) Hence it is used in weather balloons.
 (3) (a) The bond enthalpy of I_2 is higher than ICl .
 (b) Hence it is easier to break ICl bond as compared to I_2 .
 (c) Therefore ICl is more reactive than I_2 .

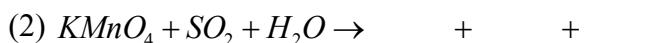
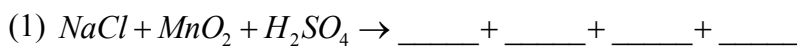


Topic: p-block element_Subtopic: Group 15, 16, 17 & 18_Level: Easy_ISC Board / Chemistry

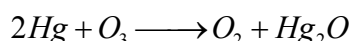
OR

- (b) (i) Explain why:
- (1) Mercury loses its meniscus in contact with ozone.
 - (2) Halogens are coloured and the colour deepens on moving down in the group from fluorine to iodine.
 - (3) Hydride of sulphur is a gas while hydride of oxygen is a liquid.

(ii) Complete and balance the following reactions :



- Ans. (i) (1) When ozone is passed through mercury, it loses its meniscus and sticks to the glass due to formation of mercurous oxide.



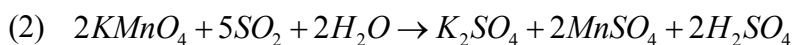
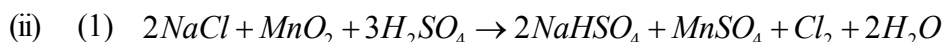
Mercurous oxide

- (2) (a) This is due to absorption of energy from visible light which results in excitation of outer electron.
 (b) As the atomic size increases from F and I , the electrons are relatively loosely bound and hence less energy is required for excitation.

(3) (a) H_2O form hydrogen bonds, hence its a liquid.

(b) In H_2S hydrogen bonding is not possible.

Topic: p-block element_Subtopic: Group 16_Level:Easy_ISC Board / Chemistry



Topic: d & f-block element_Subtopic: $KMnO_4$ reactions_Level:Easy_ISC Board / Chemistry

Question 18

(a) (i) Give balanced equations for the following reactions:

(1) Benzaldehyde reacts with hydrazine.

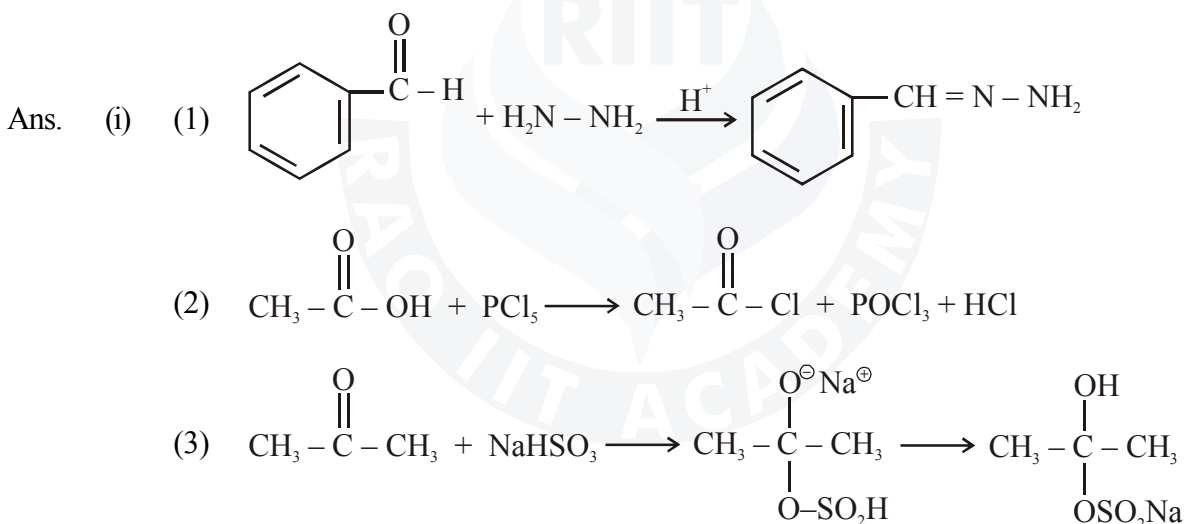
(2) Acetic acid reacts with phosphorous pentachloride.

(3) Acetone reacts with sodium bisulphite.

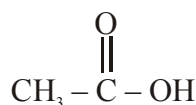
(ii) Give one chemical test each to distinguish between the following pairs of compounds:

(1) Ethanol and acetic acid

(2) Acetaldehyde and benzaldehyde



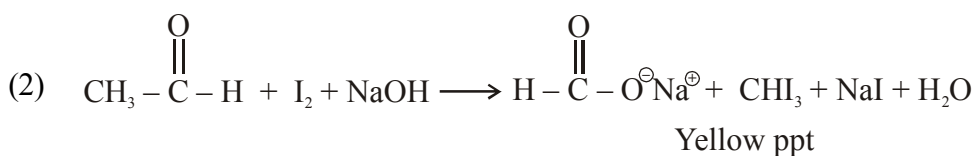
(ii) (1) $CH_3 - CH_2 - OH$

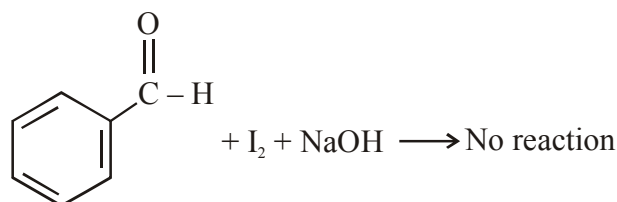


Distinguishing test between ethanol and acetic acid

Acetic acid on action of sodium bicarbonate gives CO_2 gas.

Ethanol does not react with sodium bicarbonate.





Topic: Aldehydes, ketones and carboxylic acids_Subtopic: Chemical reactions_ Level: Easy_ISC Board / Chemistry

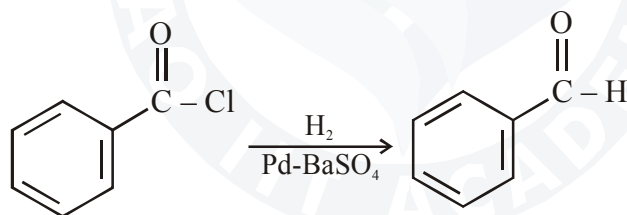
OR

- (b) (i) Write chemical equations to illustrate the following name reactions:
- (1) Clemmensen's reduction
 - (2) Rosenmund's reduction
 - (3) HVZ reaction
- (ii) Explain why:
- (1) Acetaldehyde undergoes aldol condensation, but formaldehyde does not.
 - (2) Acetic acid is weaker acid as compared to formic acid.

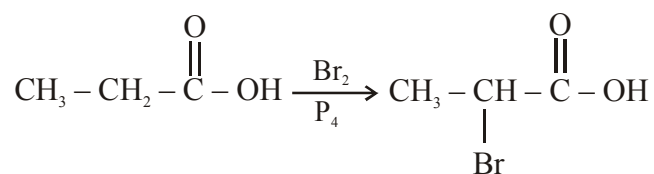
Ans. (i) (1) **Clemmensen's reduction**



(2) **Rosenmund's reduction**



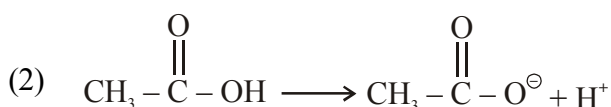
(3) **HVZ reaction**

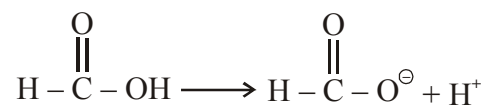


- (ii) (1) Aldol condensation is possible for aldehyde and ketone with α -hydrogen.

Acetaldehyde $\left(CH_3 - \overset{\overset{O}{\parallel}}{C} - H \right)$ has a α -hydrogen whereas formaldehyde

$\left(H - \overset{\overset{O}{\parallel}}{C} - H \right)$ does not have α -hydrogen.





Conjugate base of acetic acid, acetate ion is less stable than the conjugate base of formic acid, formate ion. This is due to the +I effect of alkyl group.

Topic: Aldehydes, ketones and carboxylic acids_Subtopic: Chemical reactions_Level: Easy_ISC Board / Chemistry

