

ICSE Board
Class X Chemistry
Board Paper – 2007

Time: 1½ hrs

Total Marks: 80

General Instructions:

1. Answers to this paper must be written on the paper provided separately.
2. You will NOT be allowed to write during the first 15 minutes. This time is to be spent in reading the question paper.

The time given at the head of this paper is the time allowed for writing the answers.

This question paper is divided into two sections.

3. **Section I** contains one question with parts (a) to (h); all the eight parts are to be answered.
4. **Section II** contains six questions numbered 2 to 7. You are to answer any four of these questions.

The intended marks of questions or for parts of questions are given in brackets [].

SECTION I (40 Marks)

Attempt **all** questions from this section.

Question 1

- (a)** From the list given below, select the word(s) required to correctly complete blanks (i) to (v) in the following passage: [5]

Ammonia, ammonium, carbonate, carbon dioxide, hydrogen, hydronium, hydroxide, precipitate, salt, water.

A solution X turns blue litmus red, so it must contain (i) _____ ions; another solution Y turns red litmus blue and therefore must contain (ii) _____ ions. When solutions X and Y are mixed together, the products will be (iii) _____ and (iv) _____. If a piece of magnesium put into solution X, (v) _____ gas would be evolved.

(Note: Words chosen from the list are to be used only once. Write the answer as (a) (i), (ii), (iii) and so on. Do not copy the passage.)

(b) From the list of characteristics given below, select the five which are relevant to non-metals and their compounds: [5]

1. Ductile
2. Conduct electricity
3. Brittle
4. Acidic oxides
5. Basic oxides
6. Discharged at anode
7. Discharged at cathode
8. Ionic
9. Covalent
10. Reaction with dilute acid yields hydrogen
11. 1, 2 or 3 valence electrons
12. 5, 6 or 7 valence electrons

(Write the five letters corresponding to the correct characteristics)

(c) A group of elements in the periodic given below (Boron is the first member of the group and Thallium is the last.). [5]

Boron

Aluminium

Gallium

Indium

Thallium

Answer the following questions in relation to the above group of elements:

1. Which element has the most metallic character?
2. Which element would be expected to have the highest electronegativity?
3. If the electronic configuration of Aluminium is 2, 8, 3, how many electrons are there in the outer shell of Thallium?
4. The atomic number of Boron is 5. Write the chemical formula of the compound formed when Boron reacts with Chlorine.
5. Will the elements in the group to the right of this Boron group be more metallic or less metallic in character? Justify your answer.

(d) Match the following: [5]

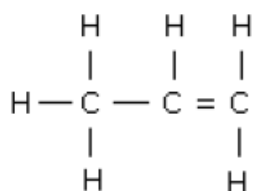
Column A

1. Acid salt
2. Mixed salt**
3. Complex salt**
4. Double salt **
5. Normal salt

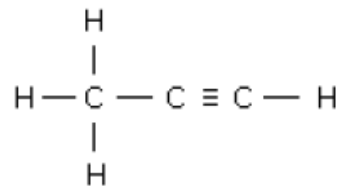
Column B

- A. Sodium potassium carbonate
- B. Alum
- C. Sodium carbonate
- D. Sodium zincate
- E. Sodium hydrogen carbonate

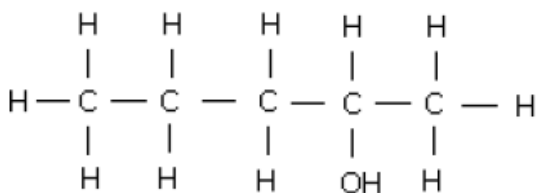
(e) Give the IUPAC names of the following compounds numbered (i) to (v). The IUPAC names of the compounds on the left are to guide you into giving the correct IUPAC name of the compounds on the right. [5]



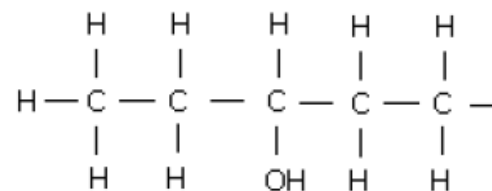
Propene



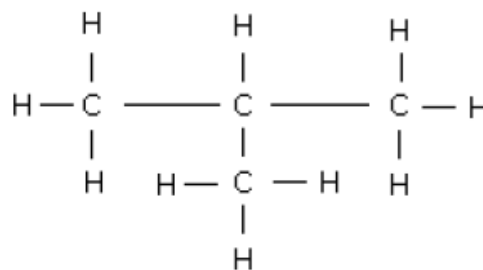
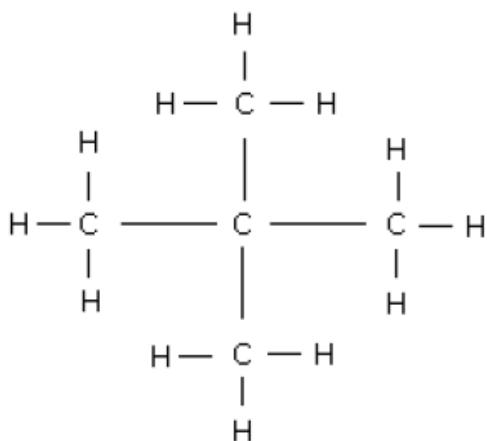
(i) _____



Pentan-2-ol

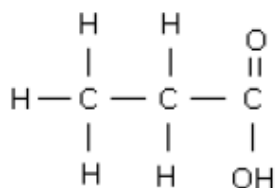


(ii) _____

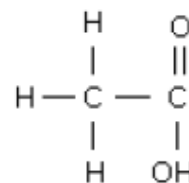


2, 2- dimethylpropane

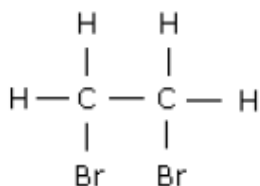
(iii) _____



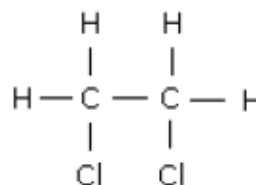
Propanoic acid



(iv) _____

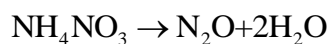


1, 2-dibromoethane



(v) _____

(f) A sample of ammonium nitrate when heated yields 8.96 litres of steam (measured at STP). [5]



1. What volume of dinitrogen oxide is produced at the same time as 8.96 litres of steam?
2. What mass of ammonium nitrate should be heated to produce 8.96 litres of steam? (Relative molecular mass of ammonium nitrate is 80)
3. Determine the percentage of oxygen in ammonium nitrate (O = 16).

(g) Write balanced equation for the following reactions: [5]

1. Lead sulphate from the lead nitrate solution and dilute sulphuric acid.
2. Copper sulphate from copper and concentrated sulphuric acid.
3. Lead chloride from lead nitrate solution and sodium chloride solution.
4. Ammonium sulphate from ammonia and dilute sulphuric acid.
5. Sodium chloride from sodium carbonate solution and dilute hydrochloric acid.

- (h)** Choose a, b, c or d to match the descriptions (i) to (v) below. Some maybe repeated.[5]
- a. Non-electrolyte
 - b. Weak electrolyte
 - c. Metallic conductor
 - d. Strong electrolyte
 - i. Molten ionic compound
 - ii. Carbon tetrachloride
 - iii. An aluminium wire
 - iv. A solution containing solvent molecules, solute molecules and ions formed by the dissociation of solute molecules.
 - v. A sugar solution with sugar molecules and water molecules.

SECTION II (40 Marks)

Attempt any **four** questions from this Section

Question 2

[10]

- (a)** Some properties of sulphuric acid are listed below. Choose the property a, b, c or d which is responsible for the reactions (i) to (v). Some properties may be repeated:

- a. Acid
 - b. Dehydrating agent
 - c. Non-volatile acid
 - d. agent
- i. $C_{12}H_{22}O_{11} + nH_2SO_4 \rightarrow 12C + 11H_2O + nH_2SO_4$
 - ii. $S + 2H_2SO_4 \rightarrow 3SO_2 + 2H_2O$
 - iii. $NaCl + H_2SO_4 \rightarrow NaHSO_4 + 2HCl$
 - iv. $CuO + H_2SO_4 \rightarrow CuSO_4 + H_2O$
 - v. $Na_2CO_3 + H_2SO_4 \rightarrow Na_2SO_4 + H_2O + CO_2$

- (b)**
1. Name the acid formed when sulphur dioxide dissolves in water.**
 2. Name the gas released when sodium carbonate is added to a solution of sulphur dioxide. **
 3. What are the two necessary conditions for the direct combination of sulphur dioxide and chlorine forming sulphuryl chloride? **
 4. State the property of sulphur dioxide which causes potassium permanganate to change its colour from purple to colourless. **

Question 3

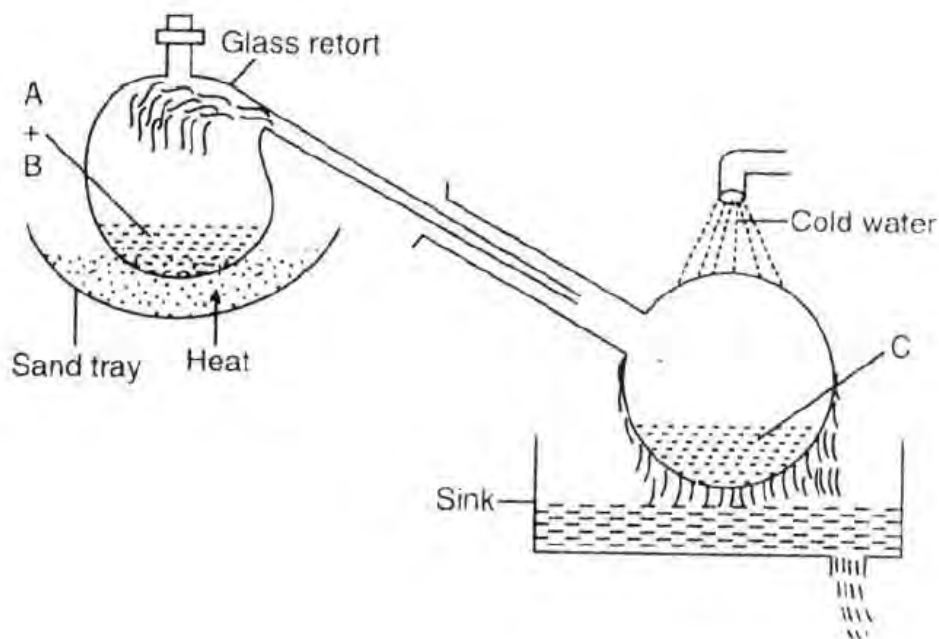
[10]

- (a)** 1. Of the two gases, ammonia and hydrogen chloride, which is more dense? Name the method of collection of this gas.
2. Give one example of a reaction between the above two gases which produces a solid compound.

(b) Write a balanced equation for a reaction in which ammonia is oxidized by:-

1. A metal oxide
2. A gas which is not oxygen

(c) The figure given below illustrates the apparatus used in the laboratory preparation of nitric acid.



1. Name A (a liquid), B (a solid) and C (a liquid) (Do not give the formulae)
2. Write an equation to show how nitric acid undergoes decomposition.
3. Write the equation for the reaction in which copper is by concentrated nitric acid.

Question 4

[10]

(a) The following is an extract from 'Metals in the Service of Man. Alexander and Street/Pelican 1976'.

"Alumina (aluminium oxide) has a very high melting point of over 2,000°C so that it cannot readily be liquefied. However, conversion of alumina to aluminium and oxygen, by electrolysis, can occur when it is dissolved in some other substance."

1. Which solution is used to react with bauxite as a first step in obtaining pure aluminium oxide?
2. The aluminium oxide for the electrolytic extraction of aluminium is obtained by heating aluminium hydroxide. Write the equation for this reaction.
3. Name the element which serves both as the anode and the cathode in the extraction of aluminium.
4. Write the equation for the reaction that occurs at the cathode during the extraction of aluminium by electrolysis.
5. Give the equation for the reaction that occurs at the anode when aluminium is purified by electrolysis.

(b) 1. Name the charged particles which attract one another to form electrovalent compound.

2. In the formation of electrovalent compounds, electrons are transferred from one element to another. How are electrons involved in the formation of a covalent compound?
3. The electronic configuration of nitrogen is 2, 5. How many electrons in the outer shell of a nitrogen atom are not involved in the formation of a nitrogen molecule?
4. In the formation of magnesium chloride (by direct combination between magnesium and chlorine), name the substance that is oxidized and the substance that is reduced.

Question 5

[10]

Copy and complete the following table which relates to three homologous series of Hydrocarbons:

General Formula	C_nH_{2n}	C_nH_{2n-2}	C_nH_{2n+2}
IUPAC name of the homologous series			
Characteristics bond type			Single bond
IUPAC name of the first member of the series			
Type of reaction with chlorine		Addition	

Question 6

[10]

- (a)** 1. HCl, HNO₃ and H₂SO₄ are the following of three compounds. Which of these compounds has the highest boiling point and which has the lowest?
2. Dilute hydrochloric acid and dilute sulphuric acid are both colourless solutions. How will the addition of barium chloride solution to each help to distinguish between the two?
3. You enter a laboratory after a class, has completed the Fountain . How will you be able to tell whether the gas used in the experiment was hydrogen chloride or ammonia?
- (b)** Write the balanced equations for the reaction of dilute hydrochloric acid with each of the following:
1. Iron
 2. Sodium hydrogen carbonate
 3. Iron (II) sulphide
 4. Sodium sulphite
 5. Sodium thiosulphate solution

Question 7

[10]

(a) A compound X consists of 4.8% carbon and 95.2% bromine by mass.

1. Determine the empirical formula of this compound working correct to one decimal place (C = 12; Br = 80).
2. If the vapour density of the compound is 252, what is the molecular formula of the compound?
3. Name the type of chemical reaction by which X can be prepared from ethane.

(b) Salts A, B, C, D and E undergo reactions (i) to (v), respectively. Identify the anion present in these salts on the basis of these reactions. Tabulate your answer in the format given below:-

1. When silver nitrate solution is added to a solution of A, a white precipitate, insoluble in dilute nitric acid, is formed.
2. Addition of dilute hydrochloric acid to B produces a gas which turns lead acetate paper black.
3. When a freshly prepared solution of ferrous sulphate is added to a solution of C and concentrated sulphuric acid is gently poured from the side of the test tube, a brown ring is formed.
4. When dilute sulphuric acid is added to D, a gas is produced which turns acidified potassium dichromate solution from orange to green.
5. Addition of dilute hydrochloric acid to E produces effervescence. The gas produced turns limewater milky but does not affect acidified potassium dichromate solution.

Salt	Anion
A	
B	
C	
D	
E	