



Set 1

CHEMISTRY

**Pre Board Examination 2017** 

Std. 12 03-01-2017

General Instructions:-

- *i)* All questions are compulsory.
- *ii)* Question numbers 1 to 5 carry one marks each.
- *iii) Question numbers 6 to 10 carry two marks each.*
- iv) Question numbers 11 to 22 carry three marks each.
- v) Question number 23 is a value based question carrying four marks.
- vi) Question numbers 24 to 26 carry five marks each.
- vii) Use log tables, if necessary.

1.	What type of stoichiometric defect is shown by AgCl?	1	
2.	Give an example of 'water in oil' and 'oil in water' emulsion.	1	
3.	Write a chemical reaction for the preparation of aspirin from phenol.		
4.	Which divalent metal ion has maximum paramagnetic character among the first transition series?		
5.	Write the structure of 1-Chloro-4-ethyl cyclohexane. 1		
6.	i) Write down the IUPAC name of the complex: $[Co(NH_3)_5Cl]^{2+}$		
	ii) Write the formula for the complex: Potassium tetrachloridonickelate(II).	2	
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7. Derive the relationship between relative lowering of vapour pressure and mole fraction of a volatile liquid.

(OR)

- i) Some liquids on mixing forms azeotropes. What are azeotropes?
- ii) What are ideal solutions? Give an example.
- 8. Give a simple chemical test to distinguish between following pairs of compounds:
  - i) Aniline and N-methylaniline
  - ii) Acetophenone and Benzophenone

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Max. Marks: 70

Time : 3 hrs.



9. Rate constant 'K' of a reaction varies with temperature 'T' according to the equation,

$$\log K = \log A - \frac{Ea}{2.303 R} \left(\frac{1}{T}\right)$$

where Ea is the activation energy. When a graph is plotted for log K vs  $\frac{1}{T}$ , a straight line with a slope of -4250 K is obtained. Calculate Ea for the reaction. R=8.314JK<sup>-1</sup>mol<sup>-1</sup>.

#### 10. What happens when

- i) PCl<sub>5</sub> is heated?
- ii)  $H_3PO_3$  is heated?

Write the reaction involved.

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- 11. Calculate the mass of a non-volatile solute (molar mass 40g/mol) which should be dissolved in 114g octane to reduce its vapour pressure to 80%. Molar Mass of Octane=114g/mol. 3
- 12. Write the names and structures of the monomers of the following polymers:
  - a) Buna-S b) Neoprene c) Nylon 6,6 3
- 13. Illustrate the following name reactions:
  - a) Cannizzaro reaction b) Clemmensen reduction
  - c) Swarts reaction
- 14. a) Write the hybridization and shape of the following complexes:

i) [CoF<sub>6</sub>]<sup>3-</sup> ii) [Ni(CN)<sub>4</sub>]<sup>2-</sup>

(Atomic number of Co=27, Ni=28)

- b) Out of NH<sub>3</sub> and CO, which ligand forms a more stable complex with a transition metal and why? 3
- 15. The edge length of unit cell of a metal having molecular mass 75g/mol is 5A°, which crystallizes in a cubic lattice. If the density is 2g/cc, then find the radius of the metal atom.

#### 16. Give reason:

- a) Ethyl iodide undergoes SN<sub>2</sub> reaction faster than ethyl bromide.
- b) 1-Chloro butane has higher boiling point than 2-Chloro butane.
- c) Grignard reagent should be prepared under anhydrous conditions.



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17. Nitrogen pentoxide decomposes according to equation :  $2N_2O_{5(g)} \rightarrow 4NO_{2(g)} + O_{2(g)}$ . This first order reaction was allowed to proceed at 40°C and the data below were collected.  $[N_2O_5](M)$  Time (min)

[N <sub>2</sub> O <sub>5</sub> ] (M)	Time (n
0.400	0.00
0.289	20.0
0.209	40.0
0.151	60.0
0.109	80.0

- a) Calculate rate constant. Include units with your answer.
- b) What will be the concentration of N<sub>2</sub>O<sub>5</sub> after 100 minutes?
- c) Calculate the initial rate of reaction.
- 18. a) What is glycogen? How is it different from starch?
  - b) Differentiate between globular and fibrous proteins.
  - c) Why Vitamin C cannot be stored in our body?
- 19. a) Which reducing agent is employed to get copper from leached low grade copper ore?
  - b) Outline the principle behind the refining of metals by the following methods.
    - i) Zone refining ii) Chromatography 3
- 20. Give reason for the following:
  - i) Helium is used in diving apparatus.
  - ii) Fluorine does not exhibit positive oxidation state.
  - iii) Oxygen shows catenation behaviour less than sulphur.

#### (OR)

Draw the structure of the following molecules:

- i) BrF<sub>3</sub> ii) (HPO<sub>3</sub>)<sub>2</sub> iii) XeF<sub>4</sub>
- How are the following conversions carried out?Benzoic acid to aniline
  - iii) Ethanamine to Propanamine.
- 22. a) In reference to Freundlich adsorption isotherm, write the expression for adsorption of gases on solids in the form of an equation.

ii)

Ethanol to propanenitrile

- b) Write an important characteristic of lyophobic sols.
- c) Write the dispersed phase and dispersion medium of smoke.
- 23. Upasana's younger brother likes taking medicines. He sometimes even drinks cough syrups even when he is not ill. One such day ,he drank cough syrup when he was healthy. After sometime he started feeling headache and his body started itching. Upasana's father did not



take him to the doctor and wanted to give medication on his own. Upasana insists that her father should not give medicines to her brother on his own but should take him to a doctor.

- Mention the values shown by Upasana. i)
- ii) Why did his body started itching?
- Why should not medicines be taken without consulting doctors? iii)
- iv) What are analgesics?
- Complete the following chemical reaction: 24. a)
  - i)
  - $\begin{array}{rcl} MnO_{4}^{-} + & C_{2}O_{4}^{2^{-}} + & H^{+} \rightarrow \\ Cr_{2}O_{7}^{2^{-}} + & Fe^{2^{+}} + & H^{+} \rightarrow \end{array}$ ii)
  - Give reason for the following: b)
    - $Mn^{2+}$  is much more resistant than  $Fe^{2+}$  towards oxidation. i)
    - Many of the transition elements are known to form interstitial compounds. ii)
    - The second and the third transition series elements have almost similar iii) atomic radii.

(OR)

Assign reason for the following:

- The enthalpies of atomization of transition elements are high. i)
- The transition metals and many of their compounds act as good catalyst. ii)
- iii) From element to element, the actinoid contraction is greater than the lanthanoid contraction.
- Scandium (Z=21) does not exhibit variable oxidation state and yet it is regarded as iv) a transition element.
- Cu(I) ion is not known in aqueous solution. v)

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- 25. Arrange the following in the increasing order of property indicated against each set. a)
  - i) Benzoic acid, 4-Nitro benzoic acid, 3,4-Dinitrobenzoic acid, 4-Methoxy benzoic acid (Acidic Strength)
  - $C_2H_5OH$ ,  $(CH_3)_2NH$ ,  $C_2H_5NH_2$  (Increasing order of boiling point) ii)
  - Complete the following reactions: b)
    - + NH<sub>2</sub>OH  $-\stackrel{H^+}{\longrightarrow}$ i)
    - $\langle \rangle CH_2 O \langle \rangle H_I \rightarrow$ ii)
    - iii)  $CH_3 CH = C (CH_3)_2 + HBr \longrightarrow$



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## (OR)

- a) Write the final products in each of the following reaction:
  - i)  $H_3C C \equiv C H \xrightarrow{Hg^{2+}, H_2SO_4}$
  - ii)  $C_6H_5CHO \frac{H_2NCONHNH_2}{2} \rightarrow$
- b) Two moles of organic compound A on treatment with a strong base gives two compounds B and C. Compound B on dehydrogenation with Cu gives A.
  Acidification of C yields Carboxylic acid D with molecular formula of CH<sub>2</sub>O<sub>2</sub>.
  Identify the compound A, B, C and D and write all chemical reactions involved. 5
- 26. a) Resistance of a conductivity cell filled with 0.1 mol/L KCl solution is 100 $\Omega$ . The resistance of the same cell when filled with 0.02 mol/L KCl solution is 520 $\Omega$ . Calculate the conductivity and molar conductivity of 0.02 mol/L KCl solution. The conductivity of 0.1 mol/L KCl solution is 1.29 x 10<sup>-2</sup>  $\Omega$ <sup>-1</sup> cm<sup>-1</sup>.
  - b) Calculate the cell potential of a hydrogen electrode in contact with a solution whose pH is 10.

### (OR)

- a) Define molar conductivity of a solution and explain how molar conductivity changes with change in concentration of solution for a weak and strong electrolyte.
- b) In the button cell widely used in watches, the following reaction takes place:

 $\begin{array}{l} {\sf Zn}(s) \,+\, {\sf Ag}_2 O_{(s)} + {\sf H}_2 O_{(l)} \,\rightarrow\, {\sf Zn}^{2+}{}_{(aq)} \,+ 2{\sf Ag}_{(s)} + \,2{\sf OH}^{-}{}_{(aq)} \\ {\sf Determine} \,\, {\sf E}^0 \,\, {\rm and} \,\, \Delta {\sf G}^0 \,\, {\rm for} \,\, {\rm the} \,\, {\rm reaction}. \end{array}$ 

 $E^{o}Zn^{2+}/Zn = -0.76 V$ ,  $E^{o}Ag^{+}/Ag = -0.80 V$ 

-X-X-X-X-X-X-

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