





CBSE Class 9 Science  
Revision Notes  
CHAPTER – 4  
STRUCTURE OF ATOM

Contribution to the Atomic theory Timeline

SCIENTIST	PROPOSED MODEL OF ATOM
<p>Democritus</p>  <p>Greek philosopher (400 B.C)</p>	 <p>Democritus (400 B.C.)</p>

- More than 2400 years ago, he named the smallest piece of matter “ATOMOS,” meaning “not to be cut.”
- Atoms were small, hard particles that were all made of the same material but were different shapes and sizes.
- Atoms were infinite in number, always moving and capable of joining together

1. Dalton’s Atomic theory:

SCIENTIST	PROPOSED ATOMIC MODEL
<p>John Dalton</p>  <p>English Chemist [proposed atomic theory in 1803]</p>	 <p>He proposed the Atomic theory of matter based on his experimental observations.</p>



- First recorded evidence that atoms exists.
- Using his theory, Dalton rationalized the various laws of chemical combination

Dalton's theory was based on the premise that the atoms of different elements could be distinguished by differences in their weights.

### Limitations

- The indivisibility of an atom was proved wrong, for, an atom can be further subdivided into protons, neutrons and electrons.
- The atoms of same element are similar in all respects, but isotopes of same element have different mass.
- Dalton's theory was based on the premise that the atoms of different elements could be distinguished by differences in their weights.

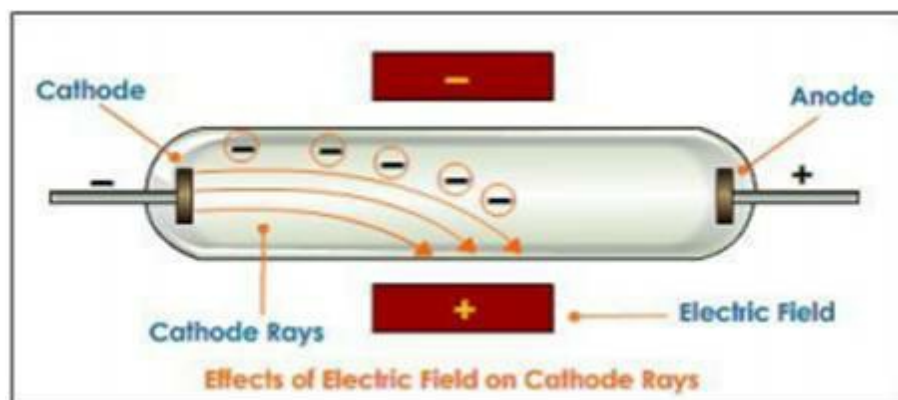
### 1. J J Thomson Experiments:

SCIENTIST	PROPOSED ATOMIC MODEL
<p>Joseph John Thomson British Physicist and Nobel laureate</p> 	<p>PLUM –PUDDING MODEL</p> 

- Discovered electrons in 1897.
- Showed us that the atom can be split into even smaller parts.

**His discovery was the first step towards a detailed model of the atom.**

- An atom is a uniform sphere of positive charges (due to presence of protons) as well as negative charges (due to presence of electrons).
- Atom as a whole is electrically neutral because the negative and positive charges are equal in magnitude.
- An electron is a negatively charged component of an atom which exists outside the nucleus. Each electron carries one unit of negative charge and has a very small mass as compared with that of a neutron or proton.




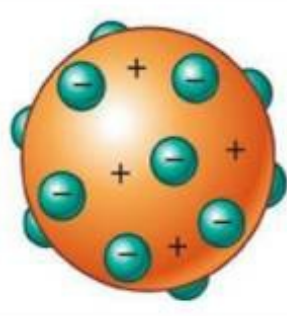
**JJ Thomson used cathode ray tubes to demonstrate that the cathode ray responds to both magnetic and electric fields.**

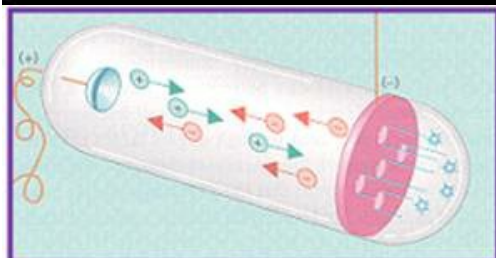
Since the ray was attracted to a positive electric plate placed over the cathode ray tube (beam deflected toward the positive plate) he determined that the ray must be composed of negatively charged particles.

He called these negative particles "electrons."

**Limitation:** Model failed to explain how protons and electrons were arranged in atom so close to each other.

**Eugene Goldstein:**

SCIENTIST	PROPOSED ATOMIC MODEL
<p>Eugene Goldstein a German physicist</p> 	





- E. Goldstein in 1886 discovered the presence of new radiations in a gas discharge and called them canal rays. These rays were positively charged radiations which ultimately led to the discovery of another sub-atomic particle.

- Used a Cathode Ray Tube to study "canal rays" which had electrical and magnetic properties opposite of an electron
- Canal Rays: The positively charged radiation produced in the discharge tube at low pressure and high voltage are called canal rays.

## Protons

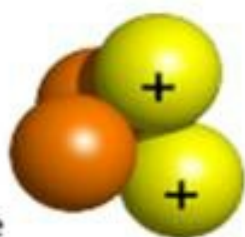
The canal rays have positively charged sub-atomic, particles known as protons (p).

### 1. Rutherford's Scattering Experiments:

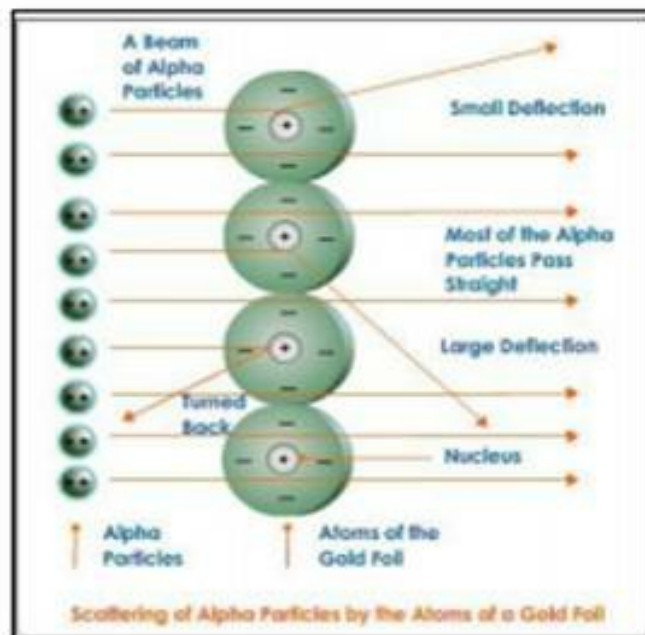
SCIENTIST	PROPOSED ATOMIC MODEL
<p>Sir Earnest Rutherford</p>  <p>Nobel prize 1908</p>	 <p>Rutherford's Nuclear Model of the Atom</p>

**Experiment:** Rutherford took a thin gold foil and made alpha particles,  $[He^{2+}]$  positively charged Helium fall on it.

S.No.	OBSERVATION	INFERENCE
1	Most of the a-particles passed through the gold foil without getting deflected.	Most of the space inside the atom is empty.
2	Very few particles were deflected.	Positive charge of the atom occupies very little space.
3	A very few alpha particles, 1 in 100000 completely rebound on hitting the goldfoil.	Nucleus of an atom is very small as compared to the total size.




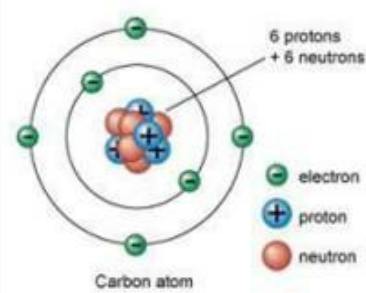
Alpha particle



- **Limitation:** In Rutherford's atomic model, Nucleus & electrons are held together by electrostatic force of attraction which would lead to the fusion between them. This does not happen in the atom.

Atomic radius  $\sim 100 \text{ pm} = 1 \times 10^{-10} \text{ m}$

Nuclear radius  $\sim 5 \times 10^{-3} \text{ pm} = 5 \times 10^{-15} \text{ m}$

SCIENTIST	PROPOSED ATOMIC MODEL
<p>James Chadwick</p>  <p>English Physicist &amp; Nobel laureate</p>	 <p>6 protons + 6 neutrons</p> <p>electron</p> <p>proton</p> <p>neutron</p> <p>Carbon atom</p>

- In 1932, James Chadwick proved that the atomic nucleus contained a neutral particle which had been proposed more than a decade earlier by Ernest Rutherford officially discovered the neutron in 1932,
- Chadwick received the Nobel Prize in 1935.

**A neutron is a subatomic particle contained in the atomic nucleus. It has no net electric, unlike the proton's positive electric charge.**

#### 4. Sub Atomic Particles:

Name	Symbol	Location in the atom	charge	Relative Mass	Actual Mass (g)
Electron	E	Around the nucleus	1-	1/1840	$9.11 \times 10^{-28}$
Proton	P	In the nucleus	1+	1	$1.67 \times 10^{-24}$
Neutron	n	In the nucleus	0	1	$1.67 \times 10^{-24}$

**Protons & Neutrons collectively are known as NUCLEONS.**

#### 5. Atomic Number & Mass Number:

**“Atomic number of an element is defined as the number of unit positive charges on the nucleus (nuclear charge) of the atom of that element or as the number of protons present in the nucleus.”**

**Atomic number, Z** = Number of unit positive charge on the nucleus

= Total number of unit positive charges carried by all protons present in the nucleus.

= Number of protons in the nucleus (p)

= Number of electrons revolving in the orbits (e)

**Eg :- Hydrogen – Atomic number = 1 (1 proton)**

**Helium - Atomic number = 2 (2 protons)**

**Mass number[ A ]** : It is defined as the sum of the number of protons & neutrons present in the nucleus of an atom.

Mass Number = Mass of protons + Mass of neutrons


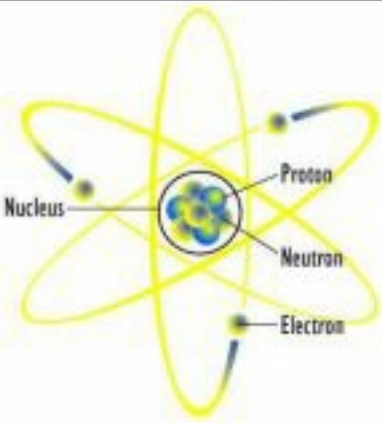
**Eg :- Carbon – Mass number = 12 (6 protons + 6 neutrons) Mass = 12u**

**Aluminium – Mass number = 27 (13 protons + 14 neutrons) Mass = 27u**

*mass number*  
*atomic number* **Symbol**

#### 6. Niel Bohr Atomic Model:



SCIENTIST	PROPOSED ATOMIC MODEL
<p>Danish physicist</p> 	

## 7. Electronic configuration & Valency:

### Bohr and Bury Scheme - Important Rules

S.No	where n = shell number	Maximum Capacity
1	K Shell $2 \times (1)$	2 electrons
2	L Shell $2 \times (2)$	8 electrons
3	M shell $2 \times (3)$	18 electrons
4	N shell $2 \times (4)$	32 electrons

**The outermost shell of an atom cannot accommodate more than 8 electrons, even if it has a capacity to accommodate more electrons. This is a very important rule and is also called the OCTET RULE. The presence of 8 electrons in the outermost shell makes the atom very stable.**

## 8. Isotopes & Isobars

### ISOTOPES

ISOTOPES	ISOBARS
Chemically same, physically different	Chemically different, physically same
Number of electrons is same	Number of electrons is different
Cannot be separated by chemical means	Can be separated by chemical means