25/6/2020
thursday

CHEMISTRY

STD - 8
class - 04

Assignment

1) Complete the table 1.17 (Page No. 26)

Ans)

Compound	Oxidation state of Mn	Subshell electronic configuration of Mn ions
$MnCl_2$	+2	$1s^2\ 2s^2\ 2p^6\ 3s^2\ 3p^6\ 3d^5$
MnO_2	+4	$1s^2\ 2s^2\ 2p^6\ 3s^2\ 3p^6\ 3d^3$
Mn_2O_3	+3	$1s^2\ 2s^2\ 2p^6\ 3s^2\ 3p^6\ 3d^4$
Mn_2O_7	+7	$1s^2\ 2s^2\ 2p^6\ 3s^2\ 3p^6$

2) Complete the table 1.16 (page no. 25)

Ans)

Compound	Oxidation state of Fe	Symbol of Fe ions
$FeCl_2$	+2	Fe^{2+}
$FeCl_3$	+3	Fe^{3+}

3) The outermost electron configuration of an element is $2s^2\ 2p^6$.

i) Which is the element?

ii) Write down the complete subshell electronic configuration.

iii) Write any two characteristics of this element.

Ans) i) Neon [Ne]

ii) $1s^2\ 2s^2\ 2p^6$

iii) * They are noble gases
* Their outermost shell is completely filled.

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