

WORKSHEET BASED ON SSLC-VICTERS CHANNEL ONLINE CLASS-PART-2-PART-3**CHEMISTRY**

1. Write the stable electronic configuration of Cr.

2. When the last electron of an atom was filled in the 3d subshell, the subshell electronic configuration was recorded as $3d^8$. Answer the questions related to this atom.

- ▶ Complete electronic Configuration
- ▶ Atomic Number
- ▶ Block
- ▶ Period number
- ▶ Group number

3. Explain the peculiarity of the electronic configurations of chromium (Cr) and copper (Cu).

4. Complete the table given below.

Element	Atomic Number	Subshell electronic configuration	Subshell to which the last electron is added.	Block
${}^3\text{Li}$ ${}^{12}\text{Mg}$ ${}^7\text{N}$ ${}^{21}\text{Sc}$				

5. Analyse the table given below and answer the following questions.

Element	A	B	C	D
Mass Number	4	23	40	16
Number of neutrons	2	12	22	8

- (a) What is the atomic number of element?
- (b) Write down the subshell electronic configuration of element D?
- (c) Which are the noble elements? why?
- (d) Arrange the elements in the increasing order of their ionization energy?

6. What is the relationship between the atomic number and the number of electrons of an element.