

24/9/2020
THURSDAY

CHEMISTRY

STD - 8
class - 17

Text book page no. 53, 54, 55

Questions and Answers.

- Which electrode has the ability to donate electrons in a cell constructed using these metals?

Ans) Zn

- Which one can gain electrons?

Ans) Cu

- Identify the chemical reaction that takes place at the Zn electrode.

Ans) $Zn \rightarrow Zn^{2+} + 2e^{-}$

- Which reaction takes place here?

Ans) Oxidation

- Write the chemical equation for the reaction taking place at the Cu electrode.

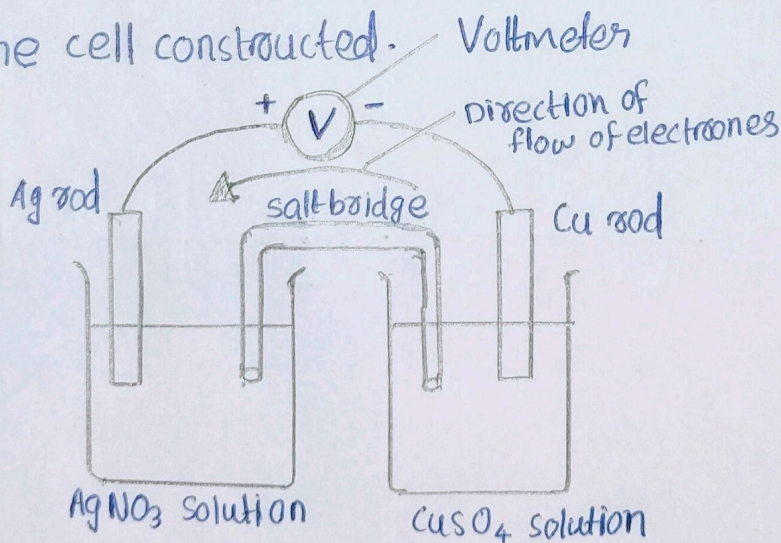
Ans) $Cu^{2+} + 2e^{-} \rightarrow Cu^0$

- What is the reaction taking place here?

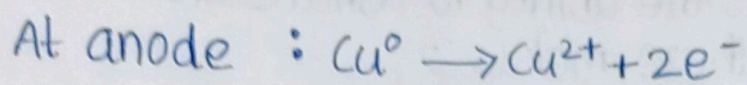
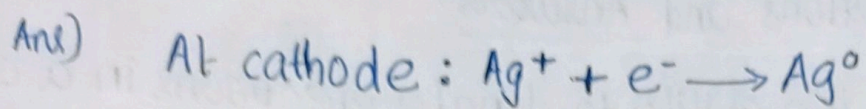
Ans) Reduction

- Sketch the cell constructed.

Ans)



- write the reactions takes place at anode and cathode.



- Complete the Table 3.4 by writing anode and cathode in each.

Ans)

Cell	Anode	Cathode
Zn-Cu	Zn	Cu
Zn-Ag	Zn	Ag
Cu-Ag	Cu	Ag