

27/10/2020
TUESDAY

CHEMISTRY

STD - 8
class - 19

Text book page no. 56, 57, 59 questions
answers.

- To which electrodes are the positive ions attracted during electrolysis?

Ans) Cathode (negative)

- To which electrodes are the negative ions attracted?

Ans) Anode (positive)

- What changes happen to the ions which are attracted to cathode?

Ans) Reduction

- What ~~happens~~ about the changes happening to the ions attracted to anode?

Ans) Oxidation

- Which ion is attracted to the positive electrode (anode)?

Ans) Cl^-

- What is the chemical reaction taking place here?

Ans) $\text{Cl}^- \rightarrow \text{Cl} + \text{e}^-$

- Which is the gas liberated at the anode?

Ans) Chlorine

- Which is the ion attracted to the negative electrode (cathode)?
Write the change happening to it?

Ans) Sodium. $\text{Na}^+ + \text{e}^- \rightarrow \text{Na}$

- Which is the metal deposited at the cathode?

Ans) Sodium

• which are the ions attracted to the positive electrode?

Ans) Cl^- , OH^-

• which are the ions attracted to the negative electrode?

Ans) Na^+ , H_3O^+

• which metal is connected to the negative terminal of the battery?

Ans) Iron bangle

• which metal is connected to the positive terminal of the battery?

Ans) Cu rod

• which solution is used as the electrolyte?

Ans) copper sulphate solution

• what happened to the copper ions?

Ans) Reduction