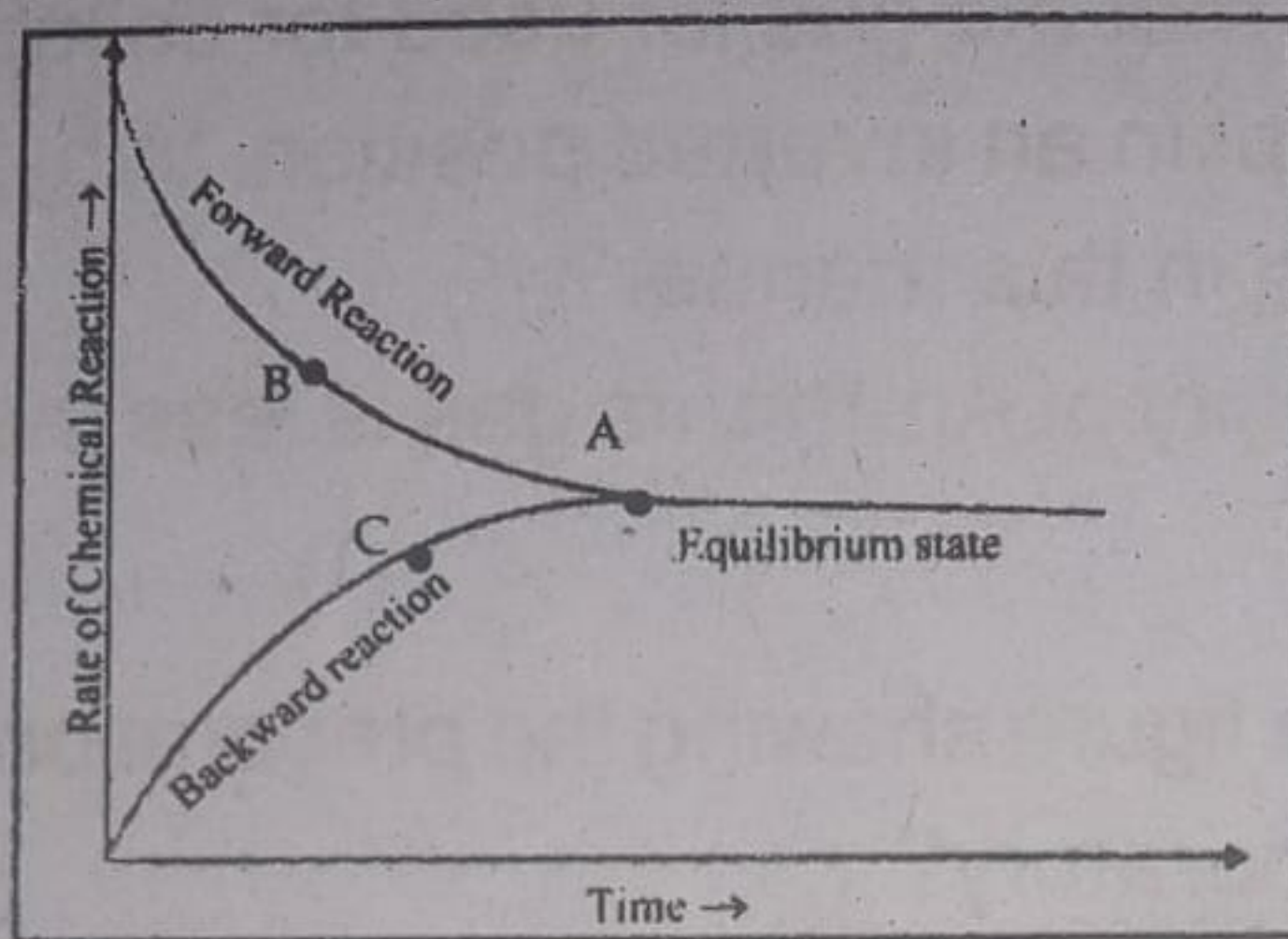



Qn. 25

The graph of a reversible reaction is given. Analyse the graph and answer the following questions ?




- As the reaction proceeds what happens to the rate of forward and backward reactions.
- At which point the rate of forward and backward reactions becomes equal ?


-  a) As reaction proceeds rate of forward reaction decreases and rate of backward reaction increases
- At the point indicated by 'A' onwards the rate of forward and backward reactions becomes equal.

Qn. 26

What happens to the rates of forward and backward reactions as time progresses?


-  As the reaction progress rates of forward reaction decreases and backward reaction increases. After sometime both becomes equal.

Qn. 28
What is chemical equilibrium ?

 In a reversible reaction when rate of forward and backward reactions become equal the system attains equilibrium. This state is called chemical equilibrium.


Qn. 29

What is a closed system?

 Closed system is one in which nothing new is added to a system or nothing is removed from the system. Equilibrium is possible only in a closed system.

Qn. 30

What are the characteristics of chemical equilibrium?

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- At the equilibrium both the reactants and the products coexist.
 - The rates of forward and backward reactions become equal at equilibrium.
 - Chemical equilibrium is dynamic at the molecular level.
 - Chemical equilibrium is attained in closed systems.