

Completion of the table

No.	Activity	Observation
1	Dropping Con. H_2SO_4 on a cotton cloth.	Cotton cloth gets burnt
2	Adding Conc. H_2SO_4 to glucose taken in a small beaker.	Black particles are formed
3	Adding Con. H_2SO_4 to a watch glass in which $CuSO_4$ crystals are taken.	<ul style="list-style-type: none">• Copper sulphate loses its colour.• The crystalline shape is lost.

Question 2.

Calcium oxide (CaO) is used as drying agent in the preparation of Ammonia in laboratory. Can concentrated H_2SO_4 be used as drying agent instead of CaO? Justify your answer.

Answer:

Sulphuric acid can not be used instead of CaO. it is because sulphuric acid reacts with ammonia and the salt ammonium sulfate is formed.

Question 3.

Which property of sulphuric acid is shown in the following situations.

a. During the preparation of chlorine, the gas is passed through concentrated H_2SO_4

b. Wooden cupboards appeared to be burnt when concentrated sulphuric acid happened to fall on it.

Answer:

a. It's character as a drying agent

b. It's character as a dehydrating agent.