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Chemistry- X- Unit -5. Class - 30

Compounds of Non - Metals

Sulphuric Acid (H_2SO_4)

Sulphuric acid is known as king of chemicals .

Uses of Sulphuric acid



Industrial production of sulphuric acid

Industrial production of sulphuric acid is known as contact process Steps

Sulphur is burnt in oxygen to produce sulphur dioxide

$$S + O_2 \rightarrow SO_2$$

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SO₂ again combines with oxygen to produce sulphur trioxide

 $\mathbf{2SO}_2 + \mathbf{O}_2 \qquad \qquad \mathbf{2SO}_3$

SO₃ is now dissolved in concentrated sulphuric acid.

 $SO_3 + H_2SO_4 \rightarrow H_2S_2O_7$

The protect thus formed is known as Oleum. Sulphuric acid is produced by dissolving Oleum in water

$H_2S_2O_7 + H_2O \rightarrow 2H_2SO_4$

Physical properties

Colourless.

Comparatively high viscosity.

Highly corrosive.

Denser than water.

Dissolves in water

Questions

1. The Industrial production of sulphuric acid in known as

(Contact process, Haber process, Ostward process)

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2. Which catalyst is used during the Industrial production of Sulphuric acid ?

3. Write any three uses of Sulphuric acid.
