

• Classification of elements into different blocks.

Group 1 & 2 \rightarrow s-block elements

Group 13 - 18 \rightarrow p-block elements

Group 3 - 12 \rightarrow d-block elements (Transition elements)

Lanthanoids and Actinoids \rightarrow f-block elements (Inner transition elements)

• General Electronic configuration of different blocks

Elements of group 3-12 constitute d-block. These are transition metals. The general electronic configuration of d-block elements is $(n-1)d^{1-10} ns^{1-2}$.

These elements show variable oxidation state.

Paramagnetism, form coloured complex ions and possess catalytic property.

Elements of group 13-18 constitute p-block. Their general electronic configuration is $ns^2 np^{1-6}$. This block include metals, non metals and metalloids.

The elements of two horizontal rows placed at the bottom of the periodic table are called f-block elements. These are also called inner transition elements. (Lanthanoids and Actinides). The general electronic configuration of f-block element is $(n-2)f^{1-14} (n-1)d^{0-1} ns^2$.