SSLC -Chemistry -Class-04

Periodic Table and Electronic Configuration

Shell and number of electrons

Shell	Shell no	Max number of electrons
K	1	2
L	2	8
M	3	18
N	4	32

Sub shells

Shell number	1	2	3	4
Subshells	s	s, p	s, p, d	s, p, d, f

Maximum number of electrons that can be accommodated in each subshell

s	2
p	6
d	10
f	14

Questions

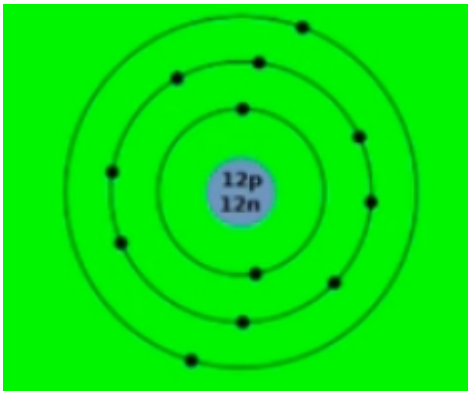
1. Maximum number of electrons that can be accommodated in N shell is.....

(2 ,8 , 18 , 32)

2. The ..... sub shell can accommodate a maximum of 10 electrons

(s , p , d , f)

3. Bohr model of Magnesium is given below . Answer the following questions.



a). What is the atomic number of this element ?

b). Write the electronic configuration of this atom ?

c). Identify the sub shells in each shell

d. How many electrons are present in each sub shell ?

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