KITE VICTERS ONLINE CLASS - 29-06-2021

SSLC -Chemistry -Class-04

Periodic Table and Electronic Configuration

Shell and number of electrons

Shell	Shell no	Max number of electrons
K	1	2
L	2	8
Μ	3	18
N	4	32

Sub shells

Shell number	1	2	3	4
Subshells	s	s, p	s, p, d	s, p, d, f

Shell (റെഷൽ)	к	L	M	N
Shell No. (n) (ഷെൽ നമ്പർ)	1	2	3	4
Subshells (സബ് ഷെല്ലുകൾ)	s	s p	s p d	s p d f

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<u>Maximum number of electrons that can be</u> <u>accommodated in each subshell</u>

S	2
р	6
d	10
f	14

Questions

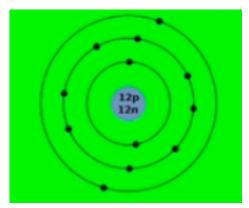
1. Maximum number of electrons that can be accommodated in N shell is.....

(2,8,18,32)

2. The sub shell can accommodate a maximum of 10 electrons

(s,p,d,f)

3. Bohr model of Magnesium is given below . Answer the following questions.



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a). What is the atomic number of this element?

b). Write the electronic configuration of this atom ?

c). Identify the sub shells in each shell

d. How many electrons are present in each sub shell ?

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