

**SSLC -Chemistry -Class -08****Periodic Table and Electronic Configuration****P Block Elements**

*p*-block

					18
13	14	15	16	17	He
B	C	N	O	F	Ne
Al	Si	P	S	Cl	Ar
Ga	Ge	As	Se	Br	Kr
In	Sn	Sb	Te	I	Xe
Tl	Pb	Bi	Po	At	Rn
Nh	Fl	Mc	Lv	Ts	Og

Groups 13 to 18 are included in the p block.

The filling of the last electron takes place in p subshell .

**Group Number of p block elements.**

p block starts in the periodic table after 12 groups.

We can find out the group number by adding 12 to the number of p electrons.

Eg:  ${}_{7}\text{N}$  ;  $1s^2 2s^2 2p^3$

Group Number :12+3=15

## **Characteristics of p block elements**

**\* The elements consists of metals ,non metals and metalloids .**

**Eg: Cl -non metal**

**Si – Metalloid**

**Al - Metal**

**\* these elements exists in three states  
Solid, Liquid and Gas .**

**Eg: C - solid , Br- Liquid , N- Gas**

**\* These elements have high ionisation energy .**

**\* They have high electronegativity values .**

**\* They have high non metallic nature .**

## **Questions**

**1. Write any two features of p block elements.**

**2. .... is p block element situated in liquid state.**

**( Mercury, Bromine ,Helium )**

**3. The element X in group 17 has 3 shells. .**

**a) Write the subshell electronic configuration of this element .?**

**b) What is the period number ?**

**c) Identify the block with which the element belong ?**

\*\*\*\*\*