

CHEMISTRY

THIRUVANANTHAPURAM EDUCATIONAL DISTRICT

AK2CH10 2(E)

WORKSHEET - 2

Standard – X

Answerkey



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- Size of the balloon increases.
 - Volume of the gas decreases with increase in pressure.
 - Boyles law states that at a constant temperature, volume of a definite mass of gas is inversely proportional to its pressure. If P is the pressure and V the volume, then $P \times V$ is a constant.
 - The size of the air bubbles rising from the bottom of an aquarium increases. (Any other relevant instance)
 - In summer the temperature is very high. According to Charles law, volume of the gas increases with increase in temperature. So during summer fully inflated tyres will burst.
 - Charles law states that at constant pressure, the volume of a definite mass of a gas is directly proportional to the temperature in Kelvin Scale. If V is volume and T the temperature, Then, V / T will be a constant.
 - 2
 - 400
 - 900
 - Volume of the balloon increases with the increase in the number of molecules present in them.
 - Avagadro's Law. It states that, at constant temperature and pressure, the volume of a gas is directly proportional to the number of molecules.
 - A = 2 litres and C = 10 litres**
 - Freedom of gas molecules is very high.
 - The energy of gas molecules is very high.
 - The attractive force between gas molecules is very less.
 - Avagadro's law
 - Charles' law
 - Boyle's law
 - Charles' law
 - A - 5**
B - 23 g
C - 12 g

CHEMISTRY

7. $B < D < A < C$

8. 1. TEMPERATURE
2. JACQUES CHARLES
3. VOLUME.
4. AMEDEO AVAGADRO
5. PRESSURE.
6. ROBERT BOYLE

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