

SSLC -Chemistry -Class -16

Unit 2 : Gas Laws and Mole Concept

Revision Questions

1) Volume of $2 \times 6.022 \times 10^{23}$ molecules of a gas at STP is

[22.4L , 44.8 L, 11.2L]

2) Which one contains $2 \times 6.022 \times 10^{23}$ Molecules ?

(28 g N_2 , 2 g H_2 , 32 g O_2 , 44.8 L CO_2)

3) 88g of CO_2 gas is taken in a container at STP.

a) Calculate the number of molecules present in the sample ?

b) Calculate the number of moles present in the sample ?

c) What is the GMM ?

d) Calculate the volume of the gas .

4) Which of the following have the same number of moles ?

[4 GMM H_2 , 88 g CO_2 , 89.6 L O_2 , 4 g He]

5) Complete the table

Substance	Volume at STP	Number of moles	Mass(g)
CO_2	44.8 L	2	88
CH_4	(a)	(b)	4 g
NH_3	11.2 L	(c)	(d)

(Hint : MM : $CO_2 = 18$, $CH_4 = 16$, $NH_3 = 17$)

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