

SSLC CHEMISTRY MODEL EXAM 2022 Answer key Prepared by Jayesh Madasseri , HMSHSS Thurakkal, Manjeri			
1	(A) S		
2	6.022×10^{23} Molecules		
3	Mg		
4	Carboxylic group		
5	NH ₃		
6	Froth floatation		
7	(B)f subshell		
8	Sodium		
9	Ions		
10	(A) a.NH ₄ Cl, Ca(OH) ₂ b.To absorb Moisture in Ammonia		
11	No of Moles in STP= $\frac{\text{Given Mass}}{\text{GMM}} = \frac{34}{17} = 2$ Mole		
12	a. CuSO ₄ b. Cu		
13	Stainless steel	Hard	For the manufacture of Utensils
	Nichrome	High Resistance	Heating coil
	Alnico	Magnetic	For Manufacture Permanent magnet
14	a.1s ² 2s ² 2p ⁶ 3s ² 3p ⁶ 4s ² b)4 c)2		
15	a. Forward reaction decreases b.Forward reaction Decreases c. Forward reaction Increases		
16	Chemical Equation		Name of reaction
	CH ₂ =CH ₂ +H ₂ ->CH ₃ -CH ₃		Addition reaction
	CH ₄ +2O ₂ --->CO ₂ +H ₂ O		Combustion
	CH ₃ - CH ₃ +Cl ₂ -->CH ₃ . CH ₂ -Cl +HCl		Substitution
17	a.(i , iv) (ii, iii) b. Alkoxy Group		
18	A)100		

	b)4 c)200 B) Boyle's law	
19	a. C ₇ H ₁₆ b. 6 c. Methyl d. 3-Methyl hexane	
20	a. Haematite b. CO c. Used as Flux to remove the impurity in the iron d. CaO+SiO ₂ ->CaSiO ₃	
21	a. Methanol b. 95.6% ethanol c. Ethanoic acid d. Glycerol	
22	a. Sugar turns into black colour b. Dehydrating nature c. If sulphuric acid reacts with Ammonia both of them lose their individual properties. d. BaCl ₂	
23	a. $13^x = 1s^2 2s^2 2p^6 3s^2 3p^1$ b. $13^x = [\text{Ne}] 3s^2 3p^1$ c. P - Block d. 7 e. $X^3 = 1s^2 2s^2 2p^6$	
24	a. Chemical energy --> electrical energy b. Zn Electrode c. Cu d. Cu e. $\text{Cu} \rightarrow \text{Cu}^{2+} + 2e^-$ (Oxidation)	