SECOND YEAR HIGHER SECONDARY EXAMINATION, MARCH 2022 PART – III CHEMISTRY MAX-MARKS – 60 SCORES

Time : 2Hrs Cool of time : 15 MNTS

1. Metal present in chlorop	gas formed by the oxidation of chloroform?	(4x1=4)
4. Which of the following (a. Propanone b. Etha	compounds answers Tollens test?	
 6. What are the ideal and n 7. Differentiate between pr 8. a) What is Arrhenius eques b) Write any 2 factors af 9. Why do transition metal 10. Write the IUPAC name a) [Co(NH₃)₄ Cl₂]Br b) K₂[HgI₄] 11. Complete the reaction a) CH₃-CH₂-Br <u>Na/dry e</u> b) Name the above react 	fecting the rate of reaction. s form coloured compounds? of following co-ordination compounds <u>ether</u>	
13. Arrange the following of HCOOH, CH ₃ COOH		
15. What is meant by denat	iline is treated with Br ₂ water? uration of proteins?	
C. Answer any 8 questions from 16 to 26. Each carries 3 score 16. Which types of deviation is shown by acetone - CHCl ₃ mixture. Draw the vapour pressure - mole fraction graph. 17. Represent the cell for the cell reaction $Zn+2 Ag^{+}_{(Aq)} \longrightarrow Zn^{2+}_{(Aq)} + 2Ag_{(S)}$ 0.1 M 0.01M		
Calculate E _{Cell}	$E_{Zn}^{0}{}^{2+}_{/Zn} = -0.76v$ $E_{Ag}^{0}{}^{+}_{/Ag} = +0.8 v$	
	pressure exerted by a solution prepared by diss $1,65000$ in 500 ml water at 27^{0} c.	olving 1g of

b) Draw the energy level diagram for crystal field splitting of octahedral complexes.

20. Write the expression for rate constant of first order reaction. Derive the expression for half life of first order reaction.

21.a)State Kohlrausch's law.

- b) Find λm^0 CH₃COONa = 91 Scm²mol⁻¹ λm^0 HCl = 426 Scm²mol⁻¹ λm^0 NaCl = 126.4 Scm²mol⁻¹
- 22. Explain the following reactions
 - a) Reimer Teimann Reaction
 - b) Williamson Synthesis
- 23. a) Why Vitamin C cannot be stored in our body?
 - b) Write any two difference between RNA and DNA
- 24. Identify the product A in the following reactions

iii)
$$C_6H_5N_2Cl \xrightarrow{H2O} A$$

25. What is Hinsberg reagent? How will you distinguish 1^{0} , 2^{0} and 3^{0} amines using this reagent?

26. Explain a) HVZ Reaction b) Gatterman - Koch reaction

D. Answer any 4 questions from 27 to 31. Each carries 4mark.

27. a) Distinguish between order and molecularity of a reaction.

b) A first order reaction is 35% complete in 40 minutes. Calculate the value of rate constants.

28. Explains the preparation of potassium dichromate from chromite ore. Write one use of $K_2Cr_2O_7$.

29. Identify the major products in the following reaction.

a) CH_3 -CH- CH_2 - CH_2 - CH_3 Alcoholic KOH heat

b) State the rule behind this reaction.

30. Explain the different types of structural isomerism in co-ordination compounds with examples.

31. How does CH₃-CHO reacts with following reagent.

a) HCN

b) diluteNaOH

- c) Zn (Hg) / Conc.HCl
- d) LiAlH₄

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